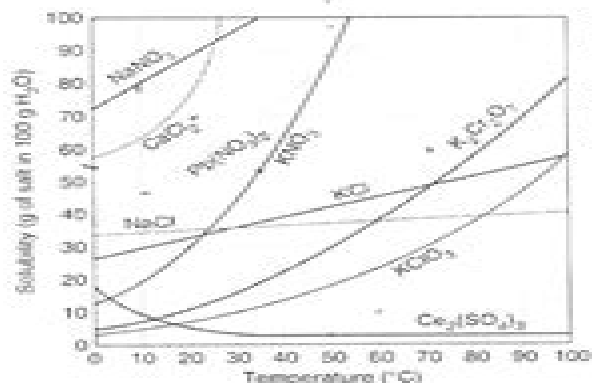


# Worksheet: Solubility Graphs

Name

Key

Use the provided solubility graph to answer the following questions:



For questions 1 - 4 an amount of solute is given, and a temperature is stated. If all of the solute could be dissolved in 100 g of water at the given temperature, would the resulting solution be unsaturated, saturated, or supersaturated?

- 60 g KCl at 70 °C Super Saturated
- 10 g KClO<sub>3</sub> at 60 °C Unsaturated
- 80 g NaNO<sub>3</sub> at 10 °C Saturated
- 70 g CaCl<sub>2</sub> at 20 °C Unsaturated

For questions 5 - 8 a solute and temperature are given. Tell how many grams of each solute must be added to 100 g of water to form a saturated solution at the given temperature.

- Pb(NO<sub>3</sub>)<sub>2</sub> at 10 °C 47g
- Ce<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub> at 50 °C 3g
- NaCl at 20 °C 35g
- K<sub>2</sub>Cr<sub>2</sub>O<sub>7</sub> at 50 °C 30g

For questions 9 and 10 underline the solution that is more concentrated.

- At 10 °C: a saturated solution of KNO<sub>3</sub> or a saturated solution of CaCl<sub>2</sub>
- At 50 °C: a saturated solution of KNO<sub>3</sub> or an unsaturated solution of NaNO<sub>3</sub> consisting of 90 g of the solute dissolved in 100 g of water.

For questions 11 - 12, show your work and circle your final answer.

- If 115 g KNO<sub>3</sub> are added to 100 g of water at 35 °C, how many grams do not dissolve? 55g = 60g will not dissolve.

- What mass of KCl would be needed to form a saturated solution if the KCl was dissolved in 200 g of water at 80 °C?

$$\frac{52g \text{ KCl}}{100g \text{ H}_2\text{O}} = \frac{104g \text{ KCl}}{200g \text{ H}_2\text{O}}$$

Same Ratio

# Progil Answers Solubility Chemistry

**Yi-Tong Ma**



## **Progil Answers Solubility Chemistry:**

## **Progil Answers Solubility Chemistry** Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become much more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Progil Answers Solubility Chemistry**," compiled by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we will delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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### **Table of Contents Progil Answers Solubility Chemistry**

1. Understanding the eBook Progil Answers Solubility Chemistry
  - The Rise of Digital Reading Progil Answers Solubility Chemistry
  - Advantages of eBooks Over Traditional Books
2. Identifying Progil Answers Solubility Chemistry
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Progil Answers Solubility Chemistry
  - User-Friendly Interface
4. Exploring eBook Recommendations from Progil Answers Solubility Chemistry
  - Personalized Recommendations
  - Progil Answers Solubility Chemistry User Reviews and Ratings
  - Progil Answers Solubility Chemistry and Bestseller Lists

5. Accessing Progil Answers Solubility Chemistry Free and Paid eBooks
  - Progil Answers Solubility Chemistry Public Domain eBooks
  - Progil Answers Solubility Chemistry eBook Subscription Services
  - Progil Answers Solubility Chemistry Budget-Friendly Options
6. Navigating Progil Answers Solubility Chemistry eBook Formats
  - ePub, PDF, MOBI, and More
  - Progil Answers Solubility Chemistry Compatibility with Devices
  - Progil Answers Solubility Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Progil Answers Solubility Chemistry
  - Highlighting and Note-Taking Progil Answers Solubility Chemistry
  - Interactive Elements Progil Answers Solubility Chemistry
8. Staying Engaged with Progil Answers Solubility Chemistry
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Progil Answers Solubility Chemistry
9. Balancing eBooks and Physical Books Progil Answers Solubility Chemistry
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Progil Answers Solubility Chemistry
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Progil Answers Solubility Chemistry
  - Setting Reading Goals Progil Answers Solubility Chemistry
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Progil Answers Solubility Chemistry
  - Fact-Checking eBook Content of Progil Answers Solubility Chemistry
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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