## BUFFERS AND THEIR PROPERTIES

## A. Determination of pH by the Use of Acid-Base Indicators

Indicator	PH Range from Observed Color with 0.10 M solution of				
Indicator	1101	NaH <sub>2</sub> PO <sub>4</sub>	HC <sub>1</sub> H <sub>1</sub> O <sub>1</sub>	ZaSO <sub>4</sub>	
Methyl violet	light blue	pumple	purple	violet	
Thymol blue	neon pink	yellow	yellow	yellow	
Methyl yellow	light pink	white gold	light pink	neon yellow	
Congo Red	Periwinkle	clear hint of	clear hint of blue	coral	
Bromocresol green	light gold yellow	lime green	light orange	blue	
PH range solution		TOTAL	STORES - STORESTS	a transport stay	

Circle the observation(s) for each solution that was most useful in estimating the pH range.

## B. Measurement of the pH of Some Typical Solutions

Record the pH and the color observed with bromocresol green for each of the 0.1 M solutions that were tested.

	NaCI	Na <sub>2</sub> CO <sub>3</sub>	NaC <sub>2</sub> H <sub>3</sub> O <sub>2</sub>	NaHSO <sub>4</sub>
рН	5.26	10.98	8.62	1.64
Color	aquamarine	blue	blue	yellow

For any two solutions having a pH less than 5 or greater than 9, write a net ionic equation to explain qualitatively why the solution has that pH. Your equation should show either a hydrogen or hydroxide ion in the products to explain the observed pH.

Solution	Net Ionic Equation		
NaHSO4	NaHSO4		
Na <sub>2</sub> CO <sub>3</sub>	Na <sub>2</sub> CO <sub>3</sub>		

# **Ph Measurement Buffers And Their Properties Asa**

**J Spring** 

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