

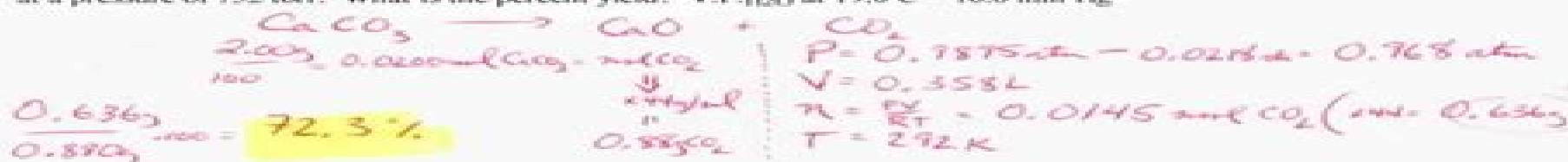
7. A student's automobile uses 1/3 gallon (1160 grams) of octane on each round trip to and from school. What volume of CO_2 gas is produced at -15°C and 29.55 in Hg?



8. Given: $2\text{NH}_4\text{Cl}_{(\text{aq})} + \text{Ca}(\text{OH})_{2(\text{aq})} \rightarrow \text{CaCl}_{2(\text{aq})} + 2\text{NH}_{3(\text{g})} + 2\text{H}_2\text{O}_{(\text{l})}$. What volume of ammonia gas is produced at 72.0°C and 3.55 atm when 6.20g of ammonium chloride reacts with 9.00g of calcium hydroxide?



9. From the decomposition of 2.00grams of CaCO_3 a student collects 358 ml of CO_2 gas over water at 19.0°C and at a pressure of 752 torr. What is the percent yield? V.P._{19.0} at 19.0°C = 16.6 mm Hg



10. 260.ml of a solution of nitric acid, HNO_3 , is analyzed with an excess of CaCO_3 by the reaction:

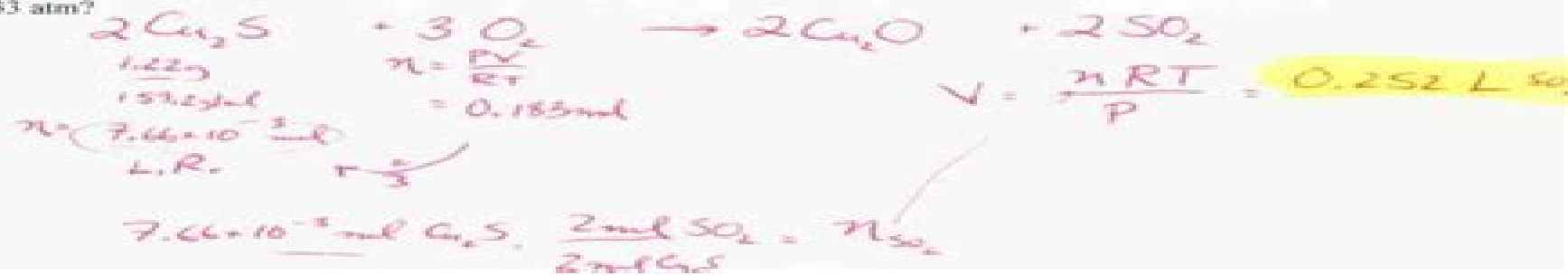


If 72.0 ml of CO_2 gas is produced at 23.0°C and 29.38 in Hg, what is the concentration of the original nitric acid in units of mg HNO_3 /ml of solution?

$$n = \frac{P \cdot V}{RT} = 2.91 \times 10^{-3}\text{mol CO}_2$$

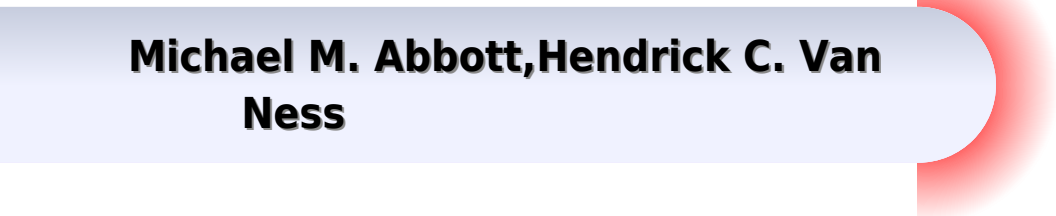


11. Consider the following unbalanced equation: $\text{Cu}_2\text{S}_{(\text{s})} + \text{O}_{2(\text{g})} \rightarrow \text{Cu}_2\text{O}_{(\text{s})} + \text{SO}_{2(\text{g})}$. If 1.22 grams of Cu_2S reacted with 4.52 liters of O_2 gas at 28°C and 1.0 atm what volume of SO_2 gas would be formed at $100.^\circ\text{C}$ and 0.933 atm?



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**Michael M. Abbott, Hendrick C. Van
Ness**



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