

## Relative Formula Mass

If  $H_2O$  is the chemical formula for water, how many atoms are in total in the substance? **3**

In chemical symbols each capital letter indicates the next element. Use a periodic table to complete the following. For the following substance list the names of the elements and the number of atoms present in the chemical formula.

$KNO_3$  \_\_\_\_\_

$Na_2CO_3$  \_\_\_\_\_

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## C2.1 - Separating Mixtures - Combined Chemistry Revision Sheet

### Pure and Impure Substances

Complete the sentences using the words in the box. Use each word once, more than once or not at all.

metals steel mineral mixed harder impure one melting many iron pure joined distilled alloys sulfur

In science a pure substance consists of \_\_\_\_\_ element/s or compound/s.

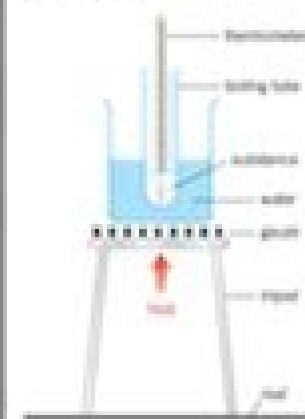
Mixtures are \_\_\_\_\_. They contain two or more different elements or

compounds \_\_\_\_\_ together \_\_\_\_\_ water and

\_\_\_\_\_ are mixtures. Mixtures of metals, called \_\_\_\_\_

are \_\_\_\_\_ than pure metals. We often determine purity of a substance

by looking at its \_\_\_\_\_ point.



Three students, A, B and C make aspirin and check its purity. The students test the purity of the aspirin by measuring the melting point, using the equipment shown. Their results are shown in the table.

What two things do they need to record to draw a change of state graph?

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The melting point of pure aspirin is in the range 138-140°C. Which student made pure aspirin?

Student	Melting point (°C)
A	140
B	139
C	137

### Filtering

Label the diagram below using the labels in the box once.

funnel filtrate beaker  
conical flask filter paper  
solution residue



Answer the following questions about when substances dissolve.

When a substance dissolves in another substance a \_\_\_\_\_ is made. A substance which

can be dissolved is said to be \_\_\_\_\_.

The substance that dissolves is called the \_\_\_\_\_.

The substance it dissolves in is called the \_\_\_\_\_.

When a substance does not dissolve it is said to be \_\_\_\_\_.

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### Empirical Formula

Empirical formula shows the simplest whole number ratio of atoms of each element in a compound.

Look at the space saving models of Butane.

Number of carbon atoms \_\_\_\_\_

Number of hydrogen atoms \_\_\_\_\_

Chemical formula \_\_\_\_\_ Empirical formula \_\_\_\_\_

Look at the displayed formula of valeric acid.

Number of carbon atoms \_\_\_\_\_

Number of hydrogen atoms \_\_\_\_\_

Number of oxygen atoms \_\_\_\_\_

Chemical formula \_\_\_\_\_ Empirical formula \_\_\_\_\_

Derive the empirical formula for the following:

$Al_2O_3$  \_\_\_\_\_  $C_4H_8O_2$  \_\_\_\_\_

$Ca_3(PO_4)_2$  \_\_\_\_\_  $Na_2SO_4$  \_\_\_\_\_

$Fe_2(SO_4)_3$  \_\_\_\_\_  $Na_2CO_3$  \_\_\_\_\_

$CaCl_2$  \_\_\_\_\_  $Na_2O$  \_\_\_\_\_

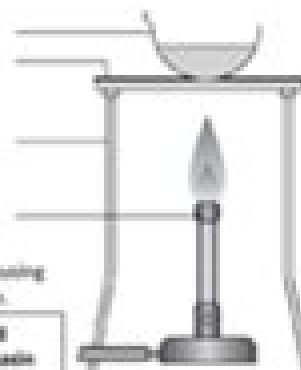
$Ca_3(PO_4)_2$  \_\_\_\_\_  $Na_2CO_3$  \_\_\_\_\_

$CaCl_2$  \_\_\_\_\_  $Na_2O$  \_\_\_\_\_

$CaCl_2$  \_\_\_\_\_  $Na_2O$  \_\_\_\_\_

### Crystallisation

Number the following steps to show the order you would carry them out to produce a sample of pure copper sulfate crystals. The first step has been done for you.



☐ Add an excess of copper carbonate powder to the acid and stir gently with a stirring rod.

☐ Carefully warm the solution of copper sulfate so that the water evaporates.

☐ Stop heating when there is a teaspoon of liquid left, allow the copper sulfate to cool and crystallise fully.

☒ Place 25 cm<sup>3</sup> sulfuric acid solution into a beaker.

☐ Fold a filter paper and place it into the filter funnel.

☐ Carefully pour the mixture through the filter paper to remove excess copper carbonate powder. The filtrate passes through the filter paper and is

\_\_\_\_\_

\_\_\_\_\_

\_\_\_\_\_

Label the diagram below using the labels in the box once.

Bunsen burner tripod  
glass evaporating basin

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**N Noddings**



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