

SECTION "A" (MULTIPLE CHOICE QUESTIONS)

1. Choose the correct answer for each from the given options: (17)

- (i) The sum of the mole fractions of solute and solvent is equal to: (a) 3 (b) 2 (c) 1 (d) 0
- (ii) The process in which a solid directly changes into vapours is known as: (a) Sublimation (b) Evaporation (c) Boiling (d) Melting
- (iii) Which one of the following is not an isotope of Hydrogen? (a) ^1_1H (b) ^2_1H (c) ^3_1H (d) ^4_1H
- (iv) The medal given to the third position holder in any event is made up of: (a) Gold (b) Silver (c) Bronze (d) Brass
- (v) In the Periodic Table, the elements have been arranged in the order of increasing: (a) Atomic Number (b) Mass Number (c) Chemical Reactivity (d) Density
- (vi) Who discovered oxygen? (a) J. Black (b) J. Priestly (c) Scheele (d) Cavendish
- (vii) Which particle is the lightest in the following? (a) Electron (b) Proton (c) Neutron (d) Alpha Particle
- (viii) Which one of the following is a liquid metal? (a) Gold (b) Silver (c) Mercury (d) Bromine
- (ix) What is the pH of a neutral solution? (a) 0 (b) 1 (c) 7 (d) 14
- (x) How many coulombs are there in one faraday? (a) 96600 (b) 96500 (c) 96650 (d) 96800
- (xi) The number of moles of solute dissolved in one litre of solution is known as: (a) Molarity (b) Molality (c) Solubility (d) Normality
- (xii) A formula that indicates the actual number and type of atoms in a molecule, is called: (a) Empirical Formula (b) Molecular Formula (c) Formula Mass (d) Molecular Mass
- (xiii) Which one of the following gases is used for respiration? (a) Oxygen (b) Hydrogen (c) Helium (d) Nitrogen

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- (xiv) Which of the following is a polar molecule? (a) O_2 (b) H_2 (c) Cl_2 (d) H_2O
- (xv) Which of the following elements is a good conductor of heat and electricity? (a) Carbon (b) hydrogen (c) Chlorine (d) Copper
- (xvi) Which of the following substances is called proton donor? (a) acids (b) bases (c) salts (d) none of them
- (xvii) The chemical formula of Sulphuric Acid is: (a) HCl (b) HNO_3 (c) HSO_4 (d) H_2SO_4

Ocr Past Paper Chemistry 2011

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