

IONISATION ENERGIES

FIRST IONISATION ENERGY

- Energy required to remove one electron from each atom in 1 mole of gaseous atoms to form 1 mole of gaseous X^+ ions
 - e.g. $Mg(g) \rightarrow Mg^+(g) + e^-$
 - $\Delta H_{1st} = +738 \text{ kJ/mol}$ (endothermic)

SECOND IONISATION ENERGY

- Energy required to remove one electron from each X^+ ion in 1 mole of gaseous ions to form 1 mole of gaseous X^{2+} ions
 - e.g. $Mg^+(g) \rightarrow Mg^{2+}(g) + e^-$
 - $\Delta H_{2nd} = +1461 \text{ kJ/mol}$ (endothermic)
 - Notice X^{2+} is higher than X^+

Ionisation energy depends on:

- The ion
 - More positive charge
 - ☐ more difficult to remove electron
- Distance between nucleus and electron
 - X^{2+} energy is higher than X^+ because the X^{2+} electron is further away from nucleus than second ion
 - ☐ attraction is less and less energy is needed to remove electron
- Shielding effect
 - Shielding electrons are complete shells of electrons between the nucleus and the electron to be removed
 - The inner shells repel the outer electrons, so the more inner shells
 - ☐ stronger shielding and less energy needed to remove electron

ELECTRON ENERGY LEVELS

- Energy levels = electronic shells
- Energy levels given numbers 1 (outermost/quantum number)
 - e.g. $n=1$ / $n=2$ etc.
- Provides 2 pieces of information
 - Total number of electrons in atom (sum of separate ionisation energies recorded)
 - Number of electron shells present in atom
- Example for Be (Beryllium):

IONISATION ENERGY	kJ/mol	
Be	900	2 electrons away to remove
Be^+	2071	
Be^{2+}	1461	2 electrons away to remove
Be^{3+}	2738	

- Interpretation (Be has 2 shells)
 - Energy level $n=1$ contains low energy electrons (less energy to remove)

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Raffaela Di Napoli



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