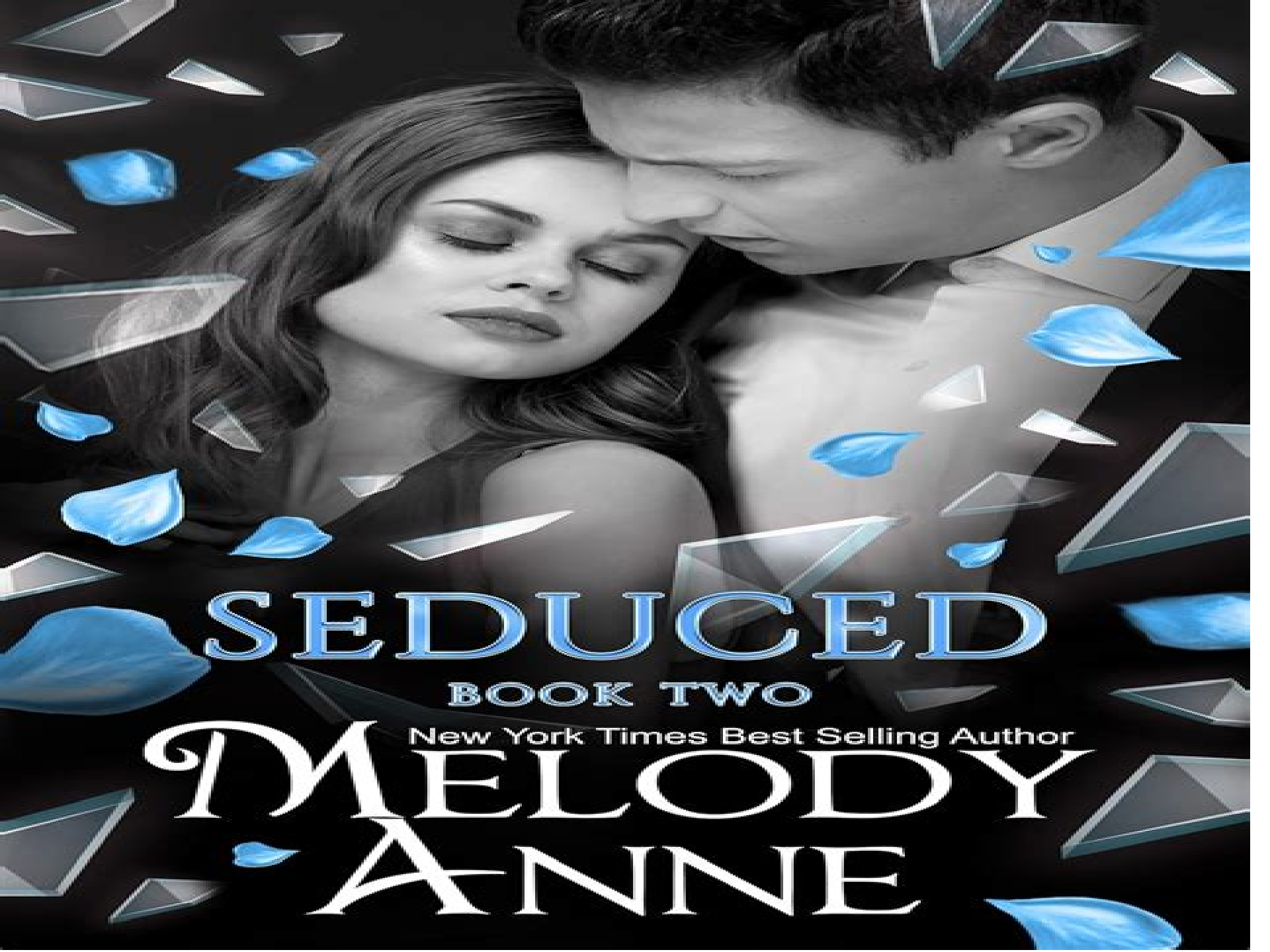


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web limiting reactant theoretical and percentage yields key $4\text{KO}_2 + 2\text{H}_2\text{O} \rightarrow 4\text{KOH} + \text{O}_2$ g if a reaction vessel contains 0 mol KO_2 and 0 mol H_2O what is the limiting reactant

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$$4\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Cu} + 2\text{Ag} + 2\text{CuNO}_3 + 2\text{AgNO}_3$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8.3 mole H₂S and multiply it with 2 mole H₂O over 2 mole H₂S from the equation 8.3 mole H₂S x 2 mole H₂O / 2 mole H₂S x 18 g H₂O = 149.4 g H₂O is your theoretical yield so percent yield = 137.1 g H₂O actual yield / 149.4 g H₂O x 100 = 91.77% problems 6 & 8 are similar to problem 5

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web a write a balanced equation for the reaction Al + O₂ → Al₂O₃ b determine the theoretical yield of Al₂O₃ 160.0 g Al x 1 mol Al / 27.0 g Al x 2 mol Al₂O₃ / 3 mol Al = 119.6 g Al₂O₃ c determine the percent yield yield actual x 100 = 260.0 g / 216.0 g x 100 = 120.4% theoretical 302.3 g

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 CH_3OH can be produced by the following reaction $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$ calculate the theoretical yield of CH_3OH if 68.5 g of CO is
 reacted with 8.6 g of H_2

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web limiting reactant and percent yield practice name 1 consider the following reaction $\text{NH}_4\text{NO}_3 + \text{Na}_3\text{PO}_4 \rightarrow \text{NH}_4\text{PO}_4 + \text{NaNO}_3$ which reactant is limiting assuming we started with 30.0 grams of ammonium nitrate and 50.0 grams of sodium phosphate what is the

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80 g I_2O_5 1 mol I_2O_5 1 mol I_2 x 333.8 g I_2O_5 1 mol I_2O_5 28 g CO 1 mol CO

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