

## 8.2

# THE NATURE OF COVALENT BONDING

# KEY

## Section Review

### Objectives

- State a rule that usually tells how many electrons are shared to form a covalent bond
- Describe how electron dot formulas are used
- Predict when two atoms are likely to be joined by a double or a triple covalent bond
- Distinguish between a single covalent bond and other covalent bonds
- Describe how the strength of a covalent bond is related to its bond dissociation energy
- Describe how resonance structures explain bonding
- Identify some exceptions to the octet rule

### Vocabulary

- single covalent bond
- structural formulas
- unshared pairs
- double covalent bonds
- triple covalent bonds
- coordinate covalent bond
- polyatomic ion
- bond dissociation energy
- resonance structures

### Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

When atoms share electrons to gain the 1 configuration of a noble gas, the bonds formed are 2. A 3 pair of valence electrons constitutes a 4 covalent bond. Pairs of valence electrons that are not shared between atoms are called 5. Sometimes two or three pairs of electrons may be shared to give 6 covalent bonds. In some cases, only one of the atoms in a bond provides the pair of bonding electrons; this is a 7. 8 is required to break covalent bonds between atoms. The total energy required to break the bond between two covalently bonded atoms is known as the 9.

When it is possible to write two or more valid electron dot formulas for a molecule or ion, each formula is referred to as a 10.

1. stable
2. covalent
3. shared
4. single
5. unshared pairs
6. double/triple
7. coordinate covalent bond
8. energy
9. bond dissociation energy
10. resonance structure

# Section 8 2 The Nature Of Covalent Bonding Worksheet

**Hongru Du**



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## Unveiling the Energy of Verbal Beauty: An Mental Sojourn through **Section 8 2 The Nature Of Covalent Bonding Worksheet**

In a world inundated with displays and the cacophony of fast communication, the profound energy and mental resonance of verbal beauty often disappear in to obscurity, eclipsed by the continuous assault of noise and distractions. However, located within the lyrical pages of **Section 8 2 The Nature Of Covalent Bonding Worksheet**, a captivating function of literary brilliance that pulses with fresh feelings, lies an unique trip waiting to be embarked upon. Composed by a virtuoso wordsmith, that magical opus books visitors on a mental odyssey, softly revealing the latent potential and profound influence embedded within the intricate internet of language. Within the heart-wrenching expanse of the evocative analysis, we shall embark upon an introspective exploration of the book is main styles, dissect their fascinating publishing fashion, and immerse ourselves in the indelible impression it leaves upon the depths of readers souls.

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