

25

NUCLEAR CHEMISTRY

Practice Problems

In your notebook, solve the following problems.

SECTION 25.1 NUCLEAR RADIATION

- What happens to the mass number and atomic number of an atom that undergoes beta decay?
- A nuclide of an element undergoes alpha particle decay. How do the atomic number and mass number of the particle change?
- Give the composition of the nucleus of the following isotopes.
 - $^{224}_{86}\text{Rn}$
 - $^{133}_{55}\text{Cs}$
 - $^{133}_{53}\text{I}$
- Complete each of the following equations.
 - $^{232}_{90}\text{Th} \rightarrow ^{228}_{88}\text{Ra} + ?$
 - $^{226}_{88}\text{Ra} \rightarrow ^{222}_{86}\text{Rn} + ?$
 - $^{234}_{90}\text{Th} \rightarrow ^{230}_{88}\text{Ra} + ?$

SECTION 25.2 NUCLEAR TRANSFORMATIONS

- Write a nuclear equation for the following radioactive processes.
 - alpha decay of francium-223
 - electron capture by beryllium-7
 - beta emission by argon-37
 - positron emission by boron-12
- Complete the equations for these transmutation reactions.
 - $^{235}_{92}\text{U} + ^1_0\text{n} \rightarrow ^{236}_{92}\text{U} + ?$
 - $^{238}_{92}\text{U} + ^1_0\text{n} \rightarrow ? + ^{140}_{54}\text{Xe} + ^{94}_{38}\text{Sr}$
 - $^{240}_{94}\text{Pu} + ^1_0\text{n} \rightarrow ? + ^{149}_{54}\text{Xe}$
 - $^{239}_{94}\text{Pu} \rightarrow ^{238}_{94}\text{Pu} + ? + ^1_0\text{n} + 4\text{ }^1_0\text{n}$
 - $^1_0\text{n} + ? \rightarrow ^{137}_{52}\text{Te} + ^{93}_{42}\text{Mo} + 2\text{ }^1_0\text{n}$
- Po-214 has a relatively short half-life of 164 s. How many micrograms would it take for 0.05 g of this isotope to decay to 0.25 g?
- How many days does it take for 10. g of gadolinium-153 to decay to 1.0 g? The half-life of gadolinium-153 is 17 days.
- By approximately what factor would the mass of a sample of copper-66 decrease in 10 minutes? The half-life of copper-66 is 5.48 min.
- In 5.48 minutes, 1.00 g of argon-39 decays to leave only 0.10 g. What is the half-life of argon-39?

Section Nuclear Radiation Review Answers Key

Edward P. Ortleb, Richard Cadice



Section Nuclear Radiation Review Answers Key:

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