

Name _____ Period _____

Acids and Bases Reinforcement Activity

Write "A" on the line if the following substance is an acid. Write "B" on the line if it is a base. If the substance is generally not an acid or a base, write "N" for neither. Most bases are metal hydroxides (OH), metal oxides (O²⁻), or ammonium (NH₄⁺) hydroxides/oxides.

1. H₂O _____

2. H₂SO₄ _____

3. NaOH _____

4. MgO _____

5. HCN _____

6. HClO₄ _____

7. HClO₃ _____

8. HClO₂ _____

9. SO₂²⁻ _____

10. HClO _____

11. NaOH _____

12. LiOH _____

13. Ba(OH)₂ _____

14. NH₄OH _____

15. Ca(OH)₂ _____

16. CaO _____

17. HNO₃ _____

18. C₆H₁₂O₆ _____

19. HNO₂ _____

20. HC₉O₈O₃ _____

21. PO₄³⁻ _____

22. Mg²⁺ _____

23. H₃C₇O₈ _____

24. H₃BO₃ _____

25. H₂SO₃ _____

26. SrO _____

27. RbOH _____

28. Rb₂O _____

29. HCl _____

30. Li₂O _____

31. HBr _____

32. NaO _____

33. HI _____

34. HF _____

35. K₂O _____

36. H₂S _____

37. C₂H₆ _____

38. C₃H₈ _____

39. KMnO₄ _____

40. H₂CO₃ _____

41. H₃PO₄ _____

42. Sr(OH)₂ _____

Complete and balance the acid-base neutralization reactions. Write "a" under the acid (reactant), and write "b" under the base (reactant). Write "c" under the salt, which is the new ionic compound produced. Hint: carbonate ions (CO₃²⁻) will split apart to produce CO₂ gas and to help form water.



Section 3 Reinforcement Acids Bases And Salts

Rosina Ehmann



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Table of Contents Section 3 Reinforcement Acids Bases And Salts

1. Understanding the eBook Section 3 Reinforcement Acids Bases And Salts
 - The Rise of Digital Reading Section 3 Reinforcement Acids Bases And Salts
 - Advantages of eBooks Over Traditional Books
2. Identifying Section 3 Reinforcement Acids Bases And Salts
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Section 3 Reinforcement Acids Bases And Salts
 - User-Friendly Interface
4. Exploring eBook Recommendations from Section 3 Reinforcement Acids Bases And Salts
 - Personalized Recommendations

- Section 3 Reinforcement Acids Bases And Salts User Reviews and Ratings
- Section 3 Reinforcement Acids Bases And Salts and Bestseller Lists
- 5. Accessing Section 3 Reinforcement Acids Bases And Salts Free and Paid eBooks
 - Section 3 Reinforcement Acids Bases And Salts Public Domain eBooks
 - Section 3 Reinforcement Acids Bases And Salts eBook Subscription Services
 - Section 3 Reinforcement Acids Bases And Salts Budget-Friendly Options
- 6. Navigating Section 3 Reinforcement Acids Bases And Salts eBook Formats
 - ePub, PDF, MOBI, and More
 - Section 3 Reinforcement Acids Bases And Salts Compatibility with Devices
 - Section 3 Reinforcement Acids Bases And Salts Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Section 3 Reinforcement Acids Bases And Salts
 - Highlighting and Note-Taking Section 3 Reinforcement Acids Bases And Salts
 - Interactive Elements Section 3 Reinforcement Acids Bases And Salts
- 8. Staying Engaged with Section 3 Reinforcement Acids Bases And Salts
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Section 3 Reinforcement Acids Bases And Salts
- 9. Balancing eBooks and Physical Books Section 3 Reinforcement Acids Bases And Salts
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Section 3 Reinforcement Acids Bases And Salts
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Section 3 Reinforcement Acids Bases And Salts
 - Setting Reading Goals Section 3 Reinforcement Acids Bases And Salts
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Section 3 Reinforcement Acids Bases And Salts
 - Fact-Checking eBook Content of Section 3 Reinforcement Acids Bases And Salts

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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