



Polarity and Intermolecular Forces

Answer Key

Vocabulary: dipole, dipole-dipole force, dipole-induced dipole force, electronegativity, intermolecular force, ionic bond, London dispersion force, molecule, nonpolar, nonpolar covalent bond, partial charges, polar, polar covalent bond, valence electron

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)



[Note: The purpose of these questions is to activate prior knowledge and get students thinking. Students are not expected to know the answers to the Prior Knowledge Questions.]

1. A big bully is having a tug-of-war with a small child. There is a ball attached to the middle of the rope.

Toward whom will the ball move? **The big bully.**

2. Two equally strong kids are having a tug-of-war. What do you expect to happen to the ball in this situation? **The ball will stay in the middle, halfway between them.**

Gizmo Warm-up

Just like in a tug-of-war, atoms that are bonded to one another pull on the electrons they share. In the Polarity and Intermolecular Forces Gizmo, you will explore how these opposing forces relate to bond types and the forces between molecules.



To begin, drag the Na (sodium) and Cl (chlorine) atoms into the simulation area. Turn on **Show valence electrons**. A **valence electron** is found in the outermost energy level of the atom.

1. Click **Play** (▶). What do you notice?

The Na atom shrinks and the Cl atom expands. The orange valence electron moves from the Na atom to the Cl atom.

2. Which atom seems to be pulling more on the sodium's one valence electron? **The Cl atom.**

How do you know? **The chlorine atom wins the "tug of war" for the valence electron.**

3. What happens to the colors of the simulated atoms, and what does this indicate?

The Na atom turns blue, indicating a positive charge. The Cl atom turns red, indicating a negative charge.



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TD Snyder

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'TD Snyder' text.

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