

11.1

DESCRIBING CHEMICAL REACTIONS

Section Review

Objectives

- 1. Represent the processes of a reaction.
- 2. Explain how to write a balanced equation.
- 3. List the steps for writing a complete ionic equation.

Vocabulary

- 1. chemical equation
- 2. balanced equation
- 3. catalyst
- 4. coefficient
- 5. skeletal equation

Part A Completion

Write the coefficient for each in the balanced equation, and balance the equation by changing the coefficient for each substance. Write balanced ionic equations for each.

1. A chemical equation is always represented by a balanced _____ 2. _____
 The chemical formula for a chemical compound is _____ 3. _____
 Chemical formulas for compounds are the _____ 4. _____
 The law of conservation of _____ 5. _____
 A chemical equation is balanced when the number of _____ 6. _____
 atoms of each element is the same on both sides of the equation. _____ 7. _____
 To balance a chemical equation, you place a _____ 8. _____
 coefficient in front of the formula of each substance. _____ 9. _____
 The equation is balanced when the number of atoms of each element is the same on both sides of the equation. _____ 10. _____

11. _____
 12. _____
 13. _____
 14. _____
 15. _____
 16. _____
 17. _____
 18. _____
 19. _____
 20. _____

1. balanced
2. chemical formula
3. chemical formula
4. law of conservation of mass
5. mass
6. coefficient
7. balanced
8. coefficient
9. balanced
10. balanced
11. _____
12. _____
13. _____
14. _____
15. _____
16. _____
17. _____
18. _____
19. _____
20. _____

Section Review 1describing Chemical Reactions

Yan Bai



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