

11.1

DESCRIBING CHEMICAL REACTIONS

Section Review

Objectives

- 1. Represent the processes of a reaction.
- 2. Explain how to write a balanced equation.
- 3. List the steps for writing a complete chemical equation.

Vocabulary

- | | |
|----------------------|----------------------|
| 1. chemical equation | 6. coefficient |
| 2. chemical equation | 7. chemical equation |
| 3. product | |

Part A Completion

Write the appropriate word(s) in space or under the blank line, and underline the word(s) in the correct blank(s). Write complete Latin names, using Roman numerals.

A chemical equation is usually represented by a symbol _____. The chemical formula for a chemical compound is _____. The law of conservation of mass requires that the _____ of the atoms in the reactants must equal the _____ of the atoms in the products. A chemical equation is balanced when the number of _____ of each element is the same on both sides of the equation. The _____ of a chemical equation is the _____ of the reaction. The _____ of a chemical equation is the _____ of the reaction.

The chemical equation for the reaction of hydrogen and oxygen to form water is _____, which is a _____ reaction. The chemical equation for the reaction of hydrogen and oxygen to form water is _____, which is a _____ reaction. The chemical equation for the reaction of hydrogen and oxygen to form water is _____, which is a _____ reaction.

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Y Pai

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular shape partially visible behind it.

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