

Module 3: Elements and the Periodic Table

Directions

Click on the following website: www.ptable.com

You will interact with this periodic table to find the answers to the questions below.

Part 1; Atomic Numbers and Atomic Masses

Complete the following table, by finding the atomic numbers and atomic masses for each element:

Symbol	P	S	Cl	Ar	K	Ca
Atomic number	15	14	17	18	19	20
Atomic Mass	30.974	28.0855	35.453	39.948	39.0983	40.078

- Using the data in the table above, is the periodic table organized by increasing atomic mass? Explain. Yes, because higher atomic numbers mean more subatomic particles which increase mass
- What property is the periodic table organized by? Atomic number

Part 2: Groups and Periods:

Continue to work in this site: www.ptable.com

Make sure the "Wikipedia" tab is selected in order to answer the following questions:

	Write your answers here:
1. The vertical columns on the periodic table are called <i>groups</i> , how many groups are there on the periodic table?	16
2. The horizontal rows on the periodic table are called <i>periods</i> , how many periods are there on the periodic table?	7
3. Which element is located in group 10 and period 5	Palladium
4. Which element is located in group 15 and period 4	Arsenic
5. Which element is located in group 2 and period 3	Magnesium
6. Which element is located in group 18 and period 6	Radon
7. Which element is located in group 1 and period 7	Francium
8. Are most of the elements on the periodic table classified as a metal or a non-metal?	Metals
9. Are the non-metal elements located on the left or right side of the periodic table?	Right

Part 3: Families

Section 3 The Periodic Table Answers

Tony Farine, Mark A. Foss



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