

Chapter 4 Atomic Structure

Section 4.2 The Structure of an Atom**(pages 108–112)**

This section compares the properties of three subatomic particles. It also discusses atomic numbers, mass numbers, and isotopes.

Reading Strategy (page 108)

Monitoring Your Understanding Before you read, list in the table shown what you know about atoms and what you would like to learn. After you read, list what you have learned. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

What I Know About Atoms	What I Would Like to Learn	What I Have Learned
Most students will know that atoms are the “building blocks” of matter, and some may know that atoms contain subatomic particles.	Based on the title of the section, students may say that they want to learn more about the structure of atoms.	

Properties of Subatomic Particles (pages 108–109)

- What are three subatomic particles?
a. Protons b. Electrons c. Neutrons
- Circle the letter that identifies a subatomic particle with a positive charge.
a. nucleus **(b)** proton
c. neutron d. electron
- Why did Chadwick conclude that the particles produced by his experiment were neutral in charge? A charged object did not deflect the paths of the particles.

Comparing Subatomic Particles (pages 109–110)

- Circle the letters of properties that vary among subatomic particles.
a. color **(b)** mass
(c) charge **(d)** location in the atom
- Circle the letter of the expression that accurately compares the masses of neutrons and protons.
(a) mass of 1 neutron = mass of 1 proton
b. mass of 2000 neutrons = mass of 1 proton
c. mass of 1 electron = mass of 1 proton
d. mass of 1 neutron = mass of 1 electron

Section Studying Atoms Answers

Michael Hoelscher

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Organic Chemistry, 12e Binder Ready Version Study Guide & Student Solutions Manual T. W. Graham Solomons, Craig B. Fryhle, Scott A. Snyder, 2016-04-11 This is the Student Study Guide Solutions Manual to accompany Organic Chemistry 12th Edition The 12th edition of Organic Chemistry continues Solomons Fryhle Snyder's tradition of excellence in teaching and preparing students for success in the organic classroom and beyond A central theme of the authors' approach to organic chemistry is to emphasize the relationship between structure and reactivity To accomplish this the content is organized in a way that combines the most useful features of a functional group approach with one largely based on reaction mechanisms The authors' philosophy is to emphasize mechanisms and their common aspects as often as possible and at the same time use the unifying features of functional groups as the basis for most chapters The structural aspects of the authors' approach show students what organic chemistry is Mechanistic aspects of their approach show students how it works And wherever an opportunity arises the authors show students what it does in living systems and the physical world around us

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