

Chapter 18 Nuclear Chemistry

Section 10.4 Fission and Fusion**(pages 308–318)**

This section discusses nuclear forces and the conversion of mass into energy. It also describes the nuclear processes of fission and fusion.

Reading Strategy (page 308)

Comparing and Contrasting As you read, contrast fission and fusion in the Venn diagram below by listing the ways they differ. For more information on this Reading Strategy, see the **Reading and Study Skills** in the Skills and Reference Handbook at the end of your textbook.

**Nuclear Forces** (pages 308–309)

- Describe the strong nuclear force. The strong nuclear force is the attractive force that binds protons and neutrons together in the nucleus.
- Is the following sentence true or false? Over very short distances, the strong nuclear force is much greater than the electric forces among protons. True
- Electric forces in atomic nuclei depend on the number of protons.
- Is the following sentence true or false? The strong nuclear force on a proton or neutron is much greater in a large nucleus than in a small nucleus. False
- All nuclei with 83 or more protons are radioactive.

Fission (pages 309–310)

- Describe fission. Fission is the splitting of an atomic nucleus into two smaller parts.
- Fission can produce very large amounts of energy from very small amounts of mass.

Section Nuclear Fission And Fusion Answer Key

Mocktime Publication



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web here s some of the guidelines for drawing dot structures so let s say we wanted to draw the dot structure for this molecule so silicon tetrafluoride the first thing we would need to do is to find the total number of valence electrons and we would account for these valence electrons in our dot structure

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web a lewis diagram shows how the valence electrons are distributed around the atoms in a molecule shared pairs of electrons are drawn as lines between atoms while lone pairs of electrons are drawn as dots next to atoms

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electrons the electron arrangement e a and the molecular geometry m g are given respectively to see a larger view of the lewis dot structure point at the molecule right click on the mouse and go to view image

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web you should try to answer the questions without referring to your textbook if you get stuck try asking another group for help for each of the following draw the lewis dot structure give the electron arrangement e a and the molecular geometry m g

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web in 1916 american chemist gilbert n lewis introduced bond lines to electron dot structures these structures also known as lewis structures or electron dot structures are drawings that visually demonstrate how electrons are shared and arranged around atoms the electrons denoted as dots are called lone pairs and belong to an individual

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chlorine molecule they share one pair of electrons

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