

Chapter 6 Chemical Bonds

Section 6.1 Ionic Bonding

(pages 155–164)

This section describes the formation of ionic bonds and the properties of ionic compounds.

Reading Strategy (page 155)

Sequencing As you read, complete the concept map to show what happens to atoms during ionic bonding. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.



Stable Electron Configurations (page 155)

- Describe the type of electron configuration that makes an atom stable and not likely to react. Atoms are stable, unreactive when they have an outer shell of eight electrons. The outer shell is called the valence shell.
- Describe an electron dot diagram. An electron dot diagram is a model in which each atom is represented by a circle with dots around it. Each dot represents a valence electron.

Ionic Bonds (pages 155–161)

- Some elements achieve stable electron configurations through the transfer of electrons between atoms.
- By losing one valence electron, a sodium atom achieves the same electron arrangement as an atom of neon.
- Circle the letter that states the result of a sodium atom transferring an electron to a chlorine atom.
 - Each atom ends up with a more stable electron arrangement.
 - The sodium atom becomes more stable, but the chlorine atom becomes less stable.
 - The chlorine atom becomes more stable, but the sodium atom becomes less stable.
 - Each atom ends up with a less stable electron arrangement.
- Is the following sentence true or false? An ion is an atom that has a net positive or negative electric charge. True
- An ion with a negative charge is called an anion.

Section 6 1 Ionic Bonding Worksheet Answer Key

Chao Zhang



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