



Samsung Intensity 111 Manual

Lauren Gardner



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web solution average atomic mass of Li 6.941 amu $\times$ 6.0151 amu $\times$ 7.0160 amu 6.941 1.0009 $\times$ 7.0160 1.0009 $\times$ 0.075 $\times$ 0.075 $\times$ 0.075 corresponds to a natural abundance of 6Li of 7.5 percent the natural abundance of 7Li is $\times$ 0.925 or 92.5 percent 3.7 the unit factor required is 6.022

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web a molar mass Li_2CO_3 2 6 941 g 12 01 g 3 16 00 g 73 89 g b molar mass CS_2 12 01 g 2 32 07 g 76 15 g c molar mass CHCl_3 12 01 g 1 008 g 3 35 45 g 119 37 g d molar mass $\text{C}_6\text{H}_8\text{O}_6$ 6 12 01 g 8 1 008 g 6 16 00 g 176 12 g e molar mass KNO_3 39 10 g 14 01 g 3 16 00 g 101 11 g

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