

Recrystallization Lab

Discussion and Conclusions

Recrystallization is a technique used for the purification of solid compounds. It is important to take into consideration that crystallization is not the same as precipitation. Precipitation is the fast formation of solid material, which may contain impurities trapped in the rapidly formed crystal lattice. On the other hand, crystallization is the slow formation of crystalline solid. Formation of crystals is a highly selective process that usually excludes foreign molecules and only crystals of the desired compound forms. Recrystallization relies on the different solubilities of solutes in a solvent. The ideal recrystallization solvent a-) should not react with the compound that is going to be purified, b-) should dissolve a moderate quantity of the substance being purified at elevated temperature, but only a small quantity at low temperature, c-) should dissolve impurities readily at low temperature or not dissolve them at all, d-) should have a boiling point below the substance's melting point, otherwise, the substance is going to melt before the solvent is removed, and e-) should be nontoxic, nonflammable and inexpensive.

In this experiment, water and 95% ethanol were tested to determine which one was the ideal solvent for unknown #15, a white solid that was selected for purification. To carry out the selection for the best solvent, 0.121 g of unknown #15 and 3 mL of water were placed in test tube #1, while 0.110 g and 3 mL of 95% ethanol were placed in test tube #2. The solubility behavior of the solute in the solvent was analyzed at room temperature, cold temperature and hot temperature. At room temperature, more than a half of the solid dissolved in 95% ethanol while only a small amount dissolved in water. Then, the test tubes were placed in an ice bath and it was observed that unknown #15 was almost insoluble in water and more than a half soluble in 95%

Recrystallization Lab Discussion

Daniel R. Palleros



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