

Unit 7 Quiz Study Guide – Lewis Structures, Polarity & VSEPR

Draw the lewis dot diagram for the following elements:

1) Na



2) Cl



3) He



4) P



5) Se



6) What are the steps for drawing a Lewis Structure?

1. Count up valence electrons – SHOW WORK
2. Determine Central Atom – C, singular/alone, more electronegativity
3. Bond terminal atoms
4. Recount
5. Fill in lone pairs outside first
6. Check if happy (STABLE OCTET. Only H, He, and B don't need 8. Everyone else NEEDS 8.)

7) Define polarity: unequal sharing of electrons in a polar bond. Polar molecules have an overall dipole. They are not symmetrical and they "move" (horse and cart analogy)

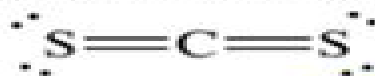
8) Define electronegativity: Atoms desire to get electrons. Fluorine is the most electronegative.

For each of the following pairs of molecules, draw the lewis structure. Then, determine which is most polar and explain your reason for making this choice:

9) carbon disulfide (CS_2)

OR

sulfur difluoride (SF_2)

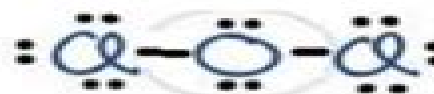


Both are nonpolar!

10) nitrogen trichloride (NCl_3)

OR

oxygen dichloride (OCl_2)



Nitrogen trichloride because it is not symmetrical and "would move"

Polarity Study Guide

J Spring



Polarity Study Guide:

A Self-study Guide to the Principles of Organic Chemistry Jiben Roy, 2013 A Self Study Guide to the Principles of Organic Chemistry Key Concepts Reaction Mechanisms and Practice Questions for the Beginner will help students new to organic chemistry grasp the key concepts of the subject quickly and easily as well as build a strong foundation for future study Starting with the definition of atom the author explains molecules electronic configuration bonding hydrocarbons polar reaction mechanisms stereochemistry reaction varieties organic spectroscopy aromaticity and aromatic reactions biomolecules organic polymers and a synthetic approach to organic compounds The over one hundred diagrams and charts contained in this volume will help students visualize the structures and bonds as they read the text and make the logic of organic chemistry clear and easily understood Each chapter ends with a list of frequently asked questions and answers followed by additional practice problems Answers are included in the Appendix Organic Chemistry, 13e Student Study Guide and Solutions Manual T. W. Graham Solomons, Craig B. Fryhle, Scott A. Snyder, 2022-05-03 Organic Chemistry Student Study Guide and Solutions Manual 13th Edition offers the full solutions for select exercises from the text **Hazmat Chemistry Study Guide (Second Edition)** Jill Meryl Levy, 2011 **ASAP Chemistry: A Quick-Review Study Guide for the AP Exam** The Princeton Review, 2019-02-12 Looking for sample exams practice questions and test taking strategies Check out our extended in depth AP chem prep guide Cracking the AP Chemistry Exam LIKE CLASS NOTES ONLY BETTER The Princeton Review s ASAP Chemistry is designed to help you zero in on just the information you need to know to successfully grapple with the AP test No questions no drills just review Advanced Placement exams require students to have a firm grasp of content you can t bluff or even logic your way to a 5 Like a set of class notes borrowed from the smartest student in your grade this book gives you exactly that No tricks or crazy stratagems no sample essays or practice sets Just the facts presented with lots of helpful visuals Inside ASAP Chemistry you ll find Essential concepts terms and functions for AP Chem all explained clearly concisely Diagrams charts and graphs for quick visual reference A three pass icon system designed to help you prioritize learning what you MUST SHOULD and COULD know in the time you have available Ask Yourself questions to help identify areas where you might need extra attention A resource that s perfect for last minute exam prep and for daily class work Topics covered in ASAP Chemistry include Atomic structure Covalent bonding intermolecular forces Thermochemistry Acids bases and more **Study Guide and Reference Material for Commercial Radio Operator Examinations** United States. Federal Communications Commission, 1955 **Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual** Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone s Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter

problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter

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