

QUESTION 3

3. 1.Learners made a mini volcano in science laboratory by adding sodium bicarbonate to ethanoic acid. They added 100 ml of a $0,2 \text{ mol}\cdot\text{dm}^{-3}$ ethanoic acid solution to 10 g of NaHCO_3 to start the reaction of the volcano.

The balanced equation for this reaction is:



- 3.1.1 Define the term limiting reagent. (2)
- 3.1.2 Determine the limiting reagent in this reaction. (6)
- 3.1.3 Calculate the mass of the other substance in excess. (3)
- 3.1.4 Calculate the volume of CO_2 produced at STP. (4)

3.2 A solution of potassium hydroxide (KOH) is prepared by dissolving 3,36 g crystals of KOH in 250 cm^3 of water.

Calculate the concentration of the potassium hydroxide solution. (4)

3.3 25 cm^3 of a potassium hydroxide solution of concentration $0,25 \text{ mol}\cdot\text{dm}^{-3}$ completely neutralises a dilute solution of sulphuric acid (H_2SO_4) in a flask.

The unbalanced equation below represents the reaction that takes place:



- 3.3.1 Balance the above equation. (2)
- 3.3.2 Calculate the mass of sulphuric acid in the flask. (5)

[25]

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1. Understanding the eBook Physical Science Exam Papers N2
 - The Rise of Digital Reading Physical Science Exam Papers N2
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Exam Papers N2
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Exam Papers N2
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Exam Papers N2
 - Personalized Recommendations
 - Physical Science Exam Papers N2 User Reviews and Ratings
 - Physical Science Exam Papers N2 and Bestseller Lists
5. Accessing Physical Science Exam Papers N2 Free and Paid eBooks
 - Physical Science Exam Papers N2 Public Domain eBooks
 - Physical Science Exam Papers N2 eBook Subscription Services
 - Physical Science Exam Papers N2 Budget-Friendly Options
6. Navigating Physical Science Exam Papers N2 eBook Formats

- ePub, PDF, MOBI, and More
- Physical Science Exam Papers N2 Compatibility with Devices
- Physical Science Exam Papers N2 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science Exam Papers N2
 - Highlighting and Note-Taking Physical Science Exam Papers N2
 - Interactive Elements Physical Science Exam Papers N2
- 8. Staying Engaged with Physical Science Exam Papers N2
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Exam Papers N2
- 9. Balancing eBooks and Physical Books Physical Science Exam Papers N2
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Exam Papers N2
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Science Exam Papers N2
 - Setting Reading Goals Physical Science Exam Papers N2
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science Exam Papers N2
 - Fact-Checking eBook Content of Physical Science Exam Papers N2
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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