

Calculus-Based Thermodynamics (Practice Problems)

1. What will be the freezing-point and boiling-point of an aqueous solution containing 100 g of glucose, $C_6H_{12}O_6$, and 100 g of water? Equations: $K_f = 1.86^\circ C/m$, $K_b = 0.512^\circ C/m$.

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} = \frac{100 \text{ g} \times \frac{1 \text{ mol}}{180 \text{ g}}}{0.100 \text{ kg}} = 5.56 \text{ mol/kg}$$

$$\Delta T_f = K_f \cdot m = 1.86^\circ C/m \times 5.56 \text{ mol/kg} = 10.3^\circ C \quad \Delta T_b = K_b \cdot m = 0.512^\circ C/m \times 5.56 \text{ mol/kg}$$

2. What will be the freezing-point and boiling-point of an aqueous solution containing 100 g of urea, CH_4N_2O , and 100 g of water? Equations: $K_f = 1.86^\circ C/m$, $K_b = 0.512^\circ C/m$. $M_{\text{urea}} = 60 \text{ g/mol}$.

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} = \frac{100 \text{ g} \times \frac{1 \text{ mol}}{60 \text{ g}}}{0.100 \text{ kg}} = 1.67 \text{ mol/kg} \quad \Delta T_f = K_f \cdot m = 1.86^\circ C/m \times 1.67 \text{ mol/kg} = 3.11^\circ C$$

$$\Delta T_b = K_b \cdot m = 0.512^\circ C/m \times 1.67 \text{ mol/kg} = 0.85^\circ C$$

3. A 100-gram sample of 0.100 M $NaCl$ is cooled to $-10.0^\circ C$. Assume that the solute does not precipitate. Equations: $K_f = 1.86^\circ C/m$. $M_{NaCl} = 58.44 \text{ g/mol}$.

$$\Delta T_f = K_f \cdot m \quad 10.0^\circ C = 1.86^\circ C/m \cdot m \quad m = 5.38 \text{ mol/kg}$$

$$\Delta T_f = K_f \cdot m \quad 10.0^\circ C = 1.86^\circ C/m \cdot m$$

$$\Rightarrow m = 5.38 \text{ mol/kg}$$

$$m = 5.38 \text{ mol/kg}$$

$$m = \frac{\text{moles solute}}{\text{kg solvent}} = \frac{100 \text{ g} \times \frac{1 \text{ mol}}{58.44 \text{ g}}}{0.100 \text{ kg}}$$

$$\Rightarrow 1.71 \text{ mol/kg} = \frac{100 \text{ g} \times \frac{1 \text{ mol}}{58.44 \text{ g}}}{0.100 \text{ kg}}$$

$$\Rightarrow 100 \text{ g} \times \frac{1 \text{ mol}}{58.44 \text{ g}} = 1.71 \text{ mol}$$

$$\Rightarrow \frac{100 \text{ g} \times \frac{1 \text{ mol}}{58.44 \text{ g}}}{0.100 \text{ kg}} = 1.71 \text{ mol/kg}$$

4. A solution composed from 10 g of glucose, $C_6H_{12}O_6$, and 100 g of water is cooled. The freezing-point of this solution is $-1.72^\circ C$. The normal freezing point of pure water is $0.00^\circ C$. What is the boiling-point of this solution? Assume that the solute does not precipitate. Equations: $K_f = 1.86^\circ C/m$, $K_b = 0.512^\circ C/m$.

$$\Delta T_f = K_f \cdot m \quad 1.72^\circ C = 1.86^\circ C/m \cdot m \quad m = 0.925 \text{ mol/kg}$$

$$\Delta T_f = K_f \cdot m \quad 1.72^\circ C = 1.86^\circ C/m \cdot m$$

$$\Rightarrow m = 0.925 \text{ mol/kg}$$

$$m = 0.925 \text{ mol/kg}$$

$$m = \frac{\text{moles solute}}{\text{kg solvent}} = \frac{10 \text{ g} \times \frac{1 \text{ mol}}{180 \text{ g}}}{0.100 \text{ kg}}$$

$$\Rightarrow 0.556 \text{ mol/kg} = \frac{10 \text{ g} \times \frac{1 \text{ mol}}{180 \text{ g}}}{0.100 \text{ kg}}$$

$$\Rightarrow 10 \text{ g} \times \frac{1 \text{ mol}}{180 \text{ g}} = 0.556 \text{ mol}$$

$$\Delta T_b = K_b \cdot m = 0.512^\circ C/m \times 0.925 \text{ mol/kg} = 0.47^\circ C$$

$$\Delta T_b = 0.47^\circ C \quad \Rightarrow \text{Boiling point} = 100.47^\circ C$$

Rosengarten Colligative Properties Homework Answers

Siddappa N.Byrareddy



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Table of Contents Rosengarten Colligative Properties Homework Answers

1. Understanding the eBook Rosengarten Colligative Properties Homework Answers
 - The Rise of Digital Reading Rosengarten Colligative Properties Homework Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Rosengarten Colligative Properties Homework Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Rosengarten Colligative Properties Homework Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Rosengarten Colligative Properties Homework Answers
 - Personalized Recommendations
 - Rosengarten Colligative Properties Homework Answers User Reviews and Ratings

- Rosengarten Colligative Properties Homework Answers and Bestseller Lists
- 5. Accessing Rosengarten Colligative Properties Homework Answers Free and Paid eBooks
 - Rosengarten Colligative Properties Homework Answers Public Domain eBooks
 - Rosengarten Colligative Properties Homework Answers eBook Subscription Services
 - Rosengarten Colligative Properties Homework Answers Budget-Friendly Options
- 6. Navigating Rosengarten Colligative Properties Homework Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Rosengarten Colligative Properties Homework Answers Compatibility with Devices
 - Rosengarten Colligative Properties Homework Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Rosengarten Colligative Properties Homework Answers
 - Highlighting and Note-Taking Rosengarten Colligative Properties Homework Answers
 - Interactive Elements Rosengarten Colligative Properties Homework Answers
- 8. Staying Engaged with Rosengarten Colligative Properties Homework Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Rosengarten Colligative Properties Homework Answers
- 9. Balancing eBooks and Physical Books Rosengarten Colligative Properties Homework Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Rosengarten Colligative Properties Homework Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Rosengarten Colligative Properties Homework Answers
 - Setting Reading Goals Rosengarten Colligative Properties Homework Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Rosengarten Colligative Properties Homework Answers
 - Fact-Checking eBook Content of Rosengarten Colligative Properties Homework Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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