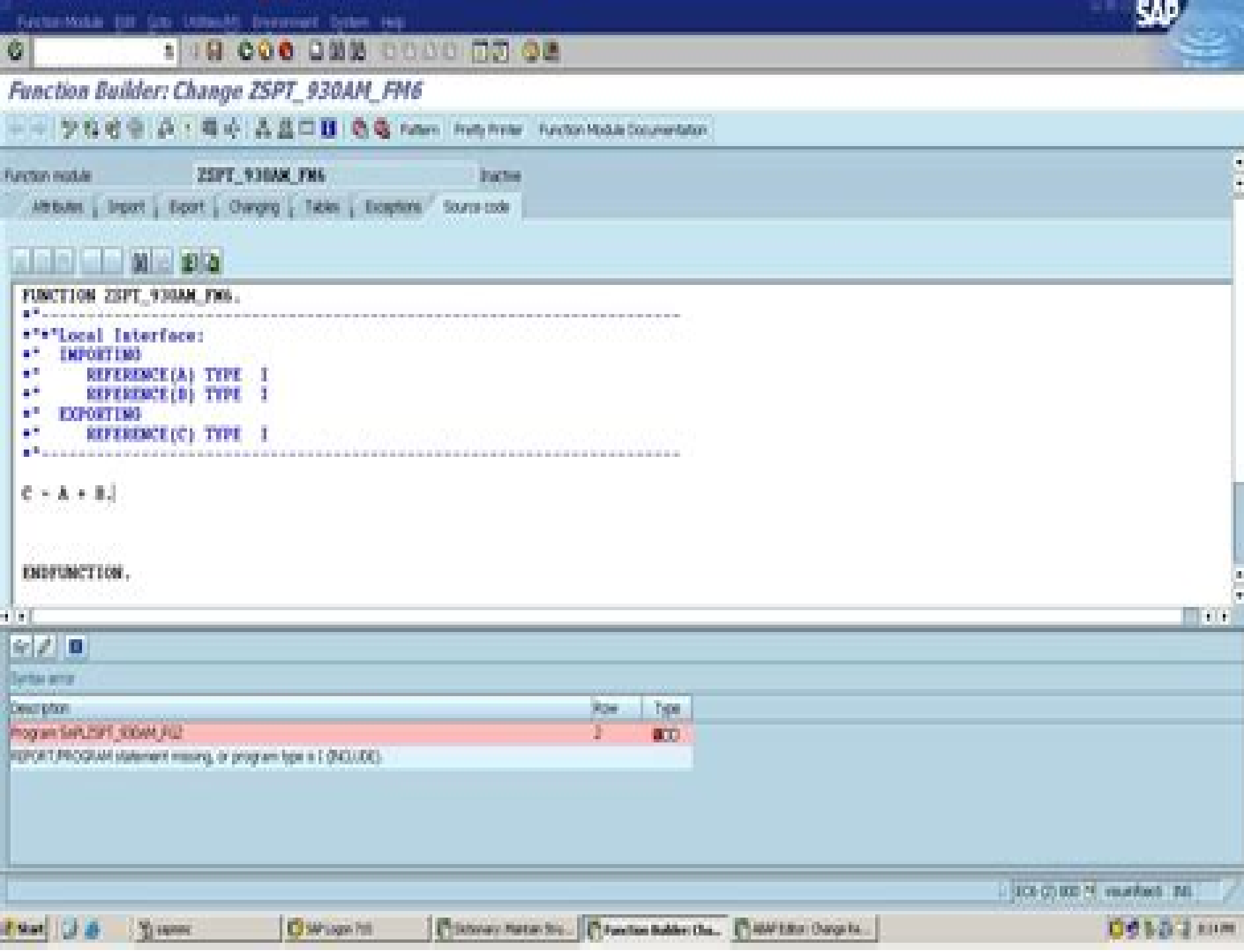




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measure an appropriate volume of each liquid eg 25 cm³ place one of the

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web for reactions involving acetic acid or ammonia the measured enthalpy change of neutralization is a few kJ less

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eg 1 HCl(aq) + NaOH(aq) → NaCl(aq) + H₂O(l) for which ΔH = 57 kJ mol⁻¹ eg 2 HNO₃(aq) + KOH(aq) → KNO₃(aq) + H₂O(l) for which ΔH = 57 kJ mol⁻¹

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