

#### 4) Polly's sensors and touch inputs

To understand how Polly works, you must know about his input devices (Fig 4).

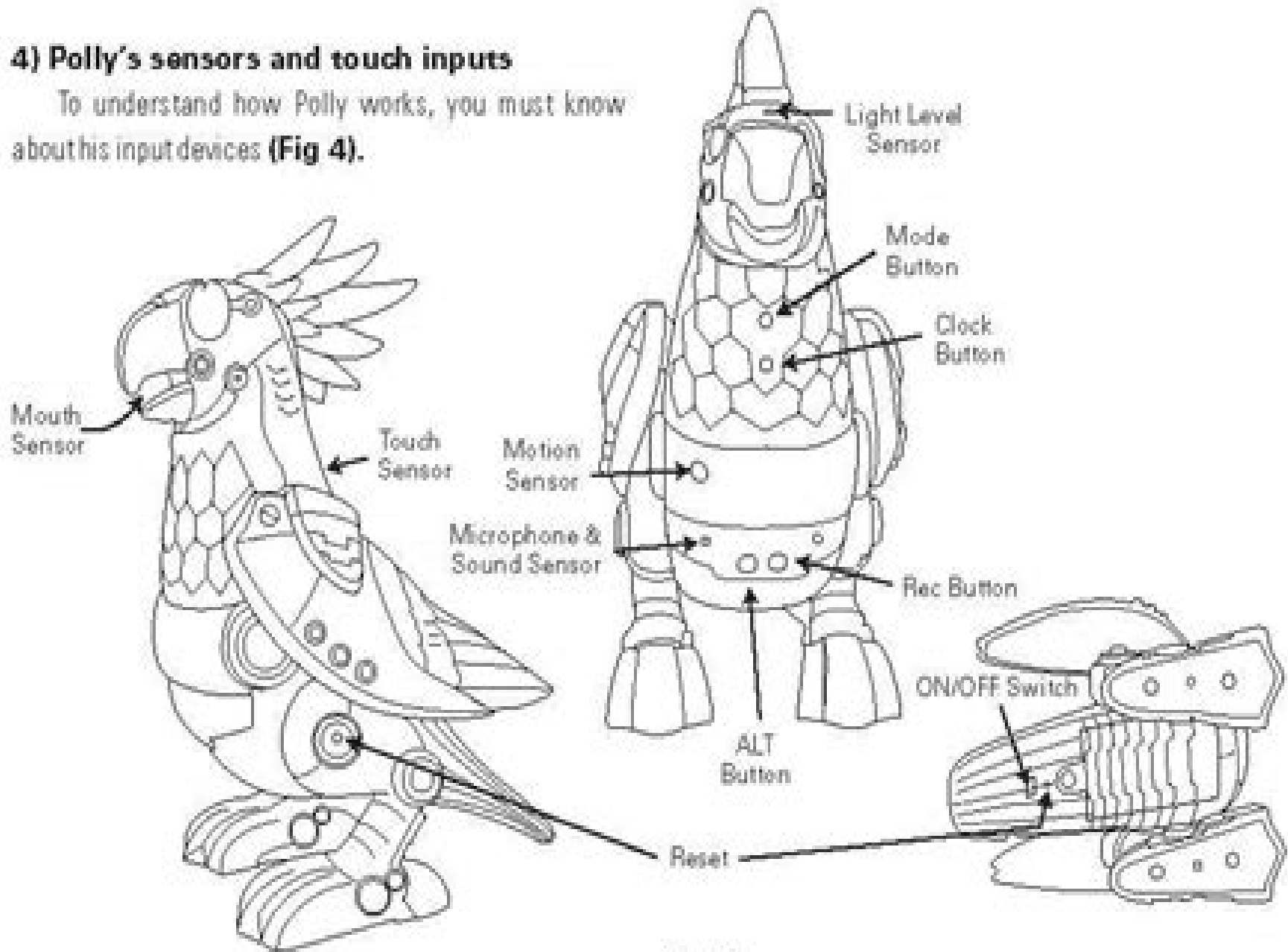


Fig. 4

# Polly The Parrot Robot Manual

**Lingjun Ying**



## **Polly The Parrot Robot Manual:**

## Unveiling the Power of Verbal Art: An Mental Sojourn through **Polly The Parrot Robot Manual**

In a world inundated with monitors and the cacophony of quick transmission, the profound energy and mental resonance of verbal beauty frequently diminish in to obscurity, eclipsed by the regular assault of sound and distractions. Yet, set within the lyrical pages of **Polly The Parrot Robot Manual**, a charming perform of literary elegance that impulses with raw thoughts, lies an memorable journey waiting to be embarked upon. Written with a virtuoso wordsmith, that mesmerizing opus instructions viewers on a mental odyssey, softly exposing the latent potential and profound impact stuck within the delicate web of language. Within the heart-wrenching expanse with this evocative evaluation, we will embark upon an introspective exploration of the book is main styles, dissect their captivating publishing design, and immerse ourselves in the indelible effect it leaves upon the depths of readers souls.

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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_2(\text{g}) + \frac{5}{2} \text{O}_2(\text{g}) \quad \Delta H = 94.5 \text{ kJ}$   
 $\text{H}_2\text{O}(\text{g}) + \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \quad \Delta H = 71.2 \text{ kJ}$   
 $\text{C}_2\text{H}_6(\text{g}) + 3 \text{H}_2\text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g}) + 1 \text{C}_2\text{H}_6(\text{g}) \quad \Delta H = 283 \text{ kJ}$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{l}) \quad \Delta H = 1411 \text{ kJ/mole}$   
 $\text{C}_2\text{H}_6(\text{g}) + 7/2 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{l}) \quad \Delta H = 1560 \text{ kJ/mole}$   
 $\text{H}_2(\text{g}) + 1/2 \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l}) \quad \Delta H = 285 \text{ kJ/mole}$

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web thus for reaction 9 2 in which 1 mole of  $h_2o$  and 1 mole of  $oh^-$  combine to form 2 moles of  $h_2o$   $\Delta h = -55.8 \text{ kJ}$  as shown note that the enthalpy change  $\Delta h_{products} - \Delta h_{reactants}$  is positive if heat is absorbed that is if  $\Delta h_{products}$

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