

Periodic Trends Worksheet

Directions: Use your notes to answer the following questions.

1. Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.

Oxygen < Carbon < Aluminum < Potassium

2. Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.

Neon < Aluminum < Sulfur < Oxygen

3. Why does fluorine have a higher ionization energy than iodine?

Fluorine has nine protons attracting 9 electrons which are much closer to the nucleus compared to iodine. This results in smaller size of Fluorine than iodine making it difficult to remove an electron from the atom.

4. Why do elements in the same family generally have similar properties?

Because they have same number of electrons in the outer shell (valence electrons) which take part in chemical reaction.

5. Indicate whether the following properties increase or decrease from left to right across the periodic table.

- | | |
|--|------------------|
| a. atomic radius (excluding noble gases) | Decreases |
| b. first ionization energy | Increases |
| c. electronegativity | Increases |

6. What trend in atomic radius occurs down a group on the periodic table? What causes this trend?

Atomic radius **increases** down the group on the periodic table. As we compare the elements down the group, the effective nuclear charge increases, but at the same time the outermost electrons are found in the shell that is farther away from the nucleus. Also the number of screening electrons increases. This results in reduced attraction between the nucleus and the outermost electrons.

7. What trend in ionization energy occurs across a period on the periodic table? What causes this trend?

Ionization energy **increases** from left to right across the period. The atomic size becomes smaller from left to right. So it becomes harder to remove electron from the atom. Hence the energy required to do so (ionization Energy) increases.

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Léo Pomar



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Table of Contents Periodic Lab Table Trend Activities Answer Key

1. Understanding the eBook Periodic Lab Table Trend Activities Answer Key
 - The Rise of Digital Reading Periodic Lab Table Trend Activities Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Lab Table Trend Activities Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Lab Table Trend Activities Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Lab Table Trend Activities Answer Key
 - Personalized Recommendations
 - Periodic Lab Table Trend Activities Answer Key User Reviews and Ratings
 - Periodic Lab Table Trend Activities Answer Key and Bestseller Lists
5. Accessing Periodic Lab Table Trend Activities Answer Key Free and Paid eBooks
 - Periodic Lab Table Trend Activities Answer Key Public Domain eBooks
 - Periodic Lab Table Trend Activities Answer Key eBook Subscription Services
 - Periodic Lab Table Trend Activities Answer Key Budget-Friendly Options
6. Navigating Periodic Lab Table Trend Activities Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Lab Table Trend Activities Answer Key Compatibility with Devices
 - Periodic Lab Table Trend Activities Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Lab Table Trend Activities Answer Key
 - Highlighting and Note-Taking Periodic Lab Table Trend Activities Answer Key
 - Interactive Elements Periodic Lab Table Trend Activities Answer Key
8. Staying Engaged with Periodic Lab Table Trend Activities Answer Key

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Periodic Lab Table Trend Activities Answer Key
- 9. Balancing eBooks and Physical Books Periodic Lab Table Trend Activities Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Lab Table Trend Activities Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Lab Table Trend Activities Answer Key
 - Setting Reading Goals Periodic Lab Table Trend Activities Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Lab Table Trend Activities Answer Key
 - Fact-Checking eBook Content of Periodic Lab Table Trend Activities Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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