

Wintersemester 2019/2020 - Klausuraufgaben 8B

Aufgabe 8 (max. 70 Minuten/100 Punkte)

Die nachstehenden Probleme sollen Sie zur Vertiefung des wissenschaftlichen Verständnisses der Stoffeigenschaften (Kapitel 10) befähigen. Sie sind, soweit in der Aufgabenstellung nicht anders angegeben, unter folgenden Bedingungen zu lösen:

$$p = p_0 = 1 \text{ bar} = 10^5 \text{ Pa}$$

1. Chloroform (Molare Masse $M = 119,38 \text{ g/mol}$) wird bei einer Temperatur von 25°C verdampft. Wie groß ist die mittlere Geschwindigkeit $\bar{v}_{\text{rms}}$?

$$\bar{v}_{\text{rms}} = \sqrt{\frac{3RT}{M}} = \sqrt{\frac{3 \cdot 8,314 \text{ J/(mol}\cdot\text{K)} \cdot 298,15 \text{ K}}{0,11938 \text{ kg/mol}}} = 316,4 \text{ m/s}$$

2. Ethylacetat (Molare Masse $M = 88,10 \text{ g/mol}$) wird bei 25°C verdampft. Wie groß ist die mittlere Geschwindigkeit $\bar{v}_{\text{rms}}$?

$$\bar{v}_{\text{rms}} = \sqrt{\frac{3RT}{M}} = \sqrt{\frac{3 \cdot 8,314 \text{ J/(mol}\cdot\text{K)} \cdot 298,15 \text{ K}}{0,08810 \text{ kg/mol}}} = 316,4 \text{ m/s}$$

3. Ethanol wird verdampft. Die mittlere Geschwindigkeit $\bar{v}_{\text{rms}}$ beträgt bei 25°C $316,4 \text{ m/s}$. Wie groß ist die mittlere Geschwindigkeit $\bar{v}_{\text{rms}}$ bei 50°C ?

$$\bar{v}_{\text{rms}} = \sqrt{\frac{3RT}{M}} = \sqrt{\frac{3 \cdot 8,314 \text{ J/(mol}\cdot\text{K)} \cdot 323,15 \text{ K}}{0,04607 \text{ kg/mol}}} = 316,4 \text{ m/s}$$

4. Die mittlere Geschwindigkeit $\bar{v}_{\text{rms}}$ beträgt bei 25°C $316,4 \text{ m/s}$. Wie groß ist die mittlere Geschwindigkeit $\bar{v}_{\text{rms}}$ bei 50°C ?

$$\bar{v}_{\text{rms}} = \sqrt{\frac{3RT}{M}} = \sqrt{\frac{3 \cdot 8,314 \text{ J/(mol}\cdot\text{K)} \cdot 323,15 \text{ K}}{0,04607 \text{ kg/mol}}} = 316,4 \text{ m/s}$$

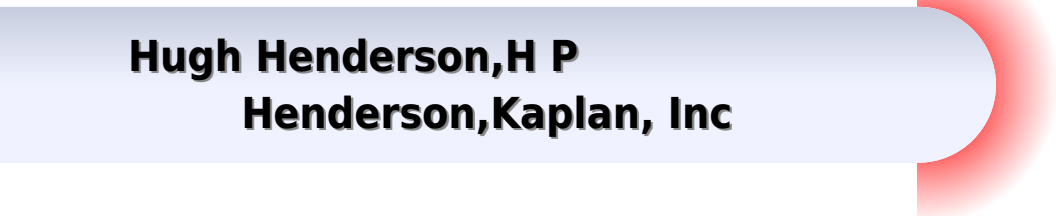
5. Ethanol wird verdampft. Die mittlere Geschwindigkeit $\bar{v}_{\text{rms}}$ beträgt bei 25°C $316,4 \text{ m/s}$. Wie groß ist die mittlere Geschwindigkeit $\bar{v}_{\text{rms}}$ bei 50°C ?

$$\bar{v}_{\text{rms}} = \sqrt{\frac{3RT}{M}} = \sqrt{\frac{3 \cdot 8,314 \text{ J/(mol}\cdot\text{K)} \cdot 323,15 \text{ K}}{0,04607 \text{ kg/mol}}} = 316,4 \text{ m/s}$$

6. Ethanol wird verdampft. Die mittlere Geschwindigkeit $\bar{v}_{\text{rms}}$ beträgt bei 25°C $316,4 \text{ m/s}$. Wie groß ist die mittlere Geschwindigkeit $\bar{v}_{\text{rms}}$ bei 50°C ?

Section 3 Gases Ch 11 Review Answers

**Hugh Henderson, H P
Henderson, Kaplan, Inc**



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