

Test Periodic Table and Trends SCH3U

Name: _____

Multiple Choice: Select the most correct answer (20K)

1. The mass number of an atom is determined by:

- a. the number of protons only
- b. adding the neutrons and protons
- c. adding the neutrons and electrons
- d. adding the protons and electrons



2. The letter "X" in the diagram above marks the:

- a. Region of greatest density
- b. Location of the protons
- c. Location of the neutrons
- d. Location of the electrons
- e. Nucleus

3. Most of the mass of the atom can be found in the:

- a. charges b. electron cloud c. electrons d. nucleus

4. Of the basic atomic particles, the one that would be attracted to a negatively charged metallic plate is the:

- a. proton b. neutron c. electron d. none of the above

5. Isotopes (such as hydrogen-1, hydrogen-2, and hydrogen-3) are atoms of the same element that differ in:

- a. the number of protons in the nucleus
- b. the number of electrons
- c. the atomic number
- d. the number of neutrons in the nucleus
- e. the mass number

6. Which of the following accurately summarizes the isotope Argon-40?

- a. 40 protons, 40 electrons, 18 neutrons
- b. 18 protons, 18 electrons, 22 neutrons
- c. 39.95 protons, 39.95 electrons, 21.05 neutrons
- d. 22 protons, 22 electrons, 18 neutrons

Periodic Trends Test Answers

Clemens Wendtner

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Periodic Trends Test Answers:

CCEA Chemistry A2 Student Unit Guide Unit 1: Periodic Trends and Further Organic, Physical and Inorganic Chemistry Alyn G. McFarland, 2014-01-31 Perfect for revision these guides explain the unit requirements summarise the content and include specimen questions with graded answers Each full colour Student Unit Guide provides ideal preparation for your unit exam Feel confident you understand the unit each guide comprehensively covers the unit content and includes topic summaries knowledge check questions and a reference index Get to grips with the exam requirements the specific skills on which you will be tested are explored and explained Analyse exam style questions graded student responses will help you focus on areas where you can improve your exam technique and performance Chemistry in a Month John Wilkinson, 2003

Chemistry for the Utterly Confused John T. Moore, Richard H. Langley, 2007-06-01 Banish bafflement in this tough subject From formulas and lab techniques to the periodic table Chemistry for the Utterly Confused focuses on the areas of maximum confusion and breaks down the most difficult chemistry topics into easy to understand concepts This invaluable guide also teaches problem solving skills you need to master this imposing subject Whether you re in high school in college or simply brushing up on chemistry knowledge this fun easily accessible book will make understanding chemistry a breeze

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