

Study Guide: Periodic Table and Periodic Trends (Honors)

1. Differentiate between Mendeleev's periodic table and Moseley's periodic table.

Mendeleev put the elements in order by atomic mass, but he noticed that in several places, the atomic masses did not always increase sequentially. Moseley discovered that ordering by atomic number was the correct way to describe the order.

2. State the periodic law.

The properties of the elements are periodic functions of their atomic numbers.

3. Vertical columns on the periodic table are called **groups or families** and horizontal rows are called **periods**.

4. Elements with similar chemical properties are placed in the same **group**.

5. Complete the table below.

| Electron Config. | Block | Period | Group/Type | Element |
|--|-------|--------|----------------------|-----------|
| [Xe] 6s ² 4f ¹⁴ 5d ¹⁰ 6p ⁶ | p | 6 | 18 / Noble Gas | Radon |
| [Xe] 6s ¹ | s | 6 | 1 / Alkali Metal | Cesium |
| [Ar] 4s ¹ 3d ⁵ | d | 4 | 6 / Trans. Metal | Chromium |
| [He] 2s ² | s | 2 | 2 / Alk. Earth Metal | Beryllium |
| [Kr] 5s ¹ 4d ¹⁰ | d | 5 | 11 / Trans. Metal | Silver |
| [He] 2s ² 2p ⁵ | p | 2 | 17 / Halogen | Fluorine |

6. Determine whether the following statements are describing alkali metals (A), alkaline earth metals (AE), transition metals (TM), halogens (H), noble gases (NG):

- a) most reactive non-metals **H**
- b) s²p⁶ valence configuration **NG**
- c) second most reactive metals **AE**
- d) s¹ valence configuration **A**
- e) exceptions found to e- configurations **TM**
- f) soft metals **A**
- g) d orbitals being filled **TM**

7. Would you expect Li and S to have different chemical properties? Why? **Yes; Li is an alkali metal with 1 valence e-, and S is a nonmetal with 6 valence e-.**

8. Explain why a magnesium atom is smaller than: a sodium atom; a calcium atom. **Mg is smaller than Na because the greater nuclear charge in Mg pulls the e- cloud in tighter (periodic trend). Mg is smaller than Ca because it has fewer energy levels (group trend).**

9. Predict the size of the astatine atom compared to the sulfur atom. **At is larger than S, because At is 3 periods down from S, but only 1 group over. The trends compete, but the sizeable increase in energy levels overcomes the slight difference in nuclear charge.**

10. Would you expect a Cl⁻ ion to be larger or smaller than an Mg²⁺ ion. Explain. **A Cl⁻ ion is larger than a Mg²⁺ ion, because the chloride ion has 3 energy levels, matching**

Periodic Trends Study Guide

Robert J. Ouellette, J. David Rawn



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Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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




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