

Unit 3 Notes: Periodic Table Notes

- John Newlands proposed an organization system based on increasing atomic mass in 1864.
- He noticed that both the chemical and physical properties repeated every 8 elements and called this the Law of Octaves.
- In 1869 both Lothar Meyer and Dmitri Mendeleev showed a connection between atomic mass and an element's properties.
- Mendeleev published first, and is given credit for this.
- He also noticed a periodic pattern when elements were ordered by increasing Atomic Mass.
- By arranging elements in order of increasing atomic mass into columns, Mendeleev created the first Periodic Table.
- This table also predicted the existence and properties of undiscovered elements.
- After many new elements were discovered, it appeared that a number of elements were out of order based on their Properties.
- In 1913 Henry Mosley discovered that each element contains a unique number of Protons.
- By rearranging the elements based on Atomic Number, the problems with the Periodic Table were corrected.
- This new arrangement creates a periodic repetition of both physical and chemical properties known as the Periodic Law.

PERIODS

A simplified periodic table diagram where horizontal rows are highlighted with different colors to represent periods. The colors from top to bottom are: red/orange (1st period), yellow (2nd period), green (3rd period), blue (4th period), purple (5th period), and pink (6th period). The lanthanide and actinide series are shown as separate rows of purple and pink boxes below the main body.

Periods are the Rows

Valence electrons across a period are in the same energy level

GROUPS

A simplified periodic table diagram where vertical columns are highlighted with different colors to represent groups. The colors from left to right are: red/orange (1st group), yellow (2nd group), green (3rd group), blue (4th group), purple (5th group), and pink (6th group). The central and bottom sections of the table are shown as empty white boxes.

Groups/Families are the Columns

There are equal numbers of valence electrons in a group.

- When elements are arranged in order of increasing Atomic Number, there is a periodic repetition of their physical and chemical properties

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