

1. Circle the element with the largest atomic radius and put a square around the element with the smallest atomic radius:

Cu K Ni Br

- a. Explain why you made these choices: All of the elements are in the same period. The trend in atomic radius as you go across a period is DECREASING. Therefore, the element on the far left (K) is the largest, and the element on the far right (Br) is the smallest.

2. Circle the element with the highest ionization energy and put a square around the element with the lowest ionization energy:

Cu K Ni Br

- a. Explain why you made these choices: All of the elements are in the same period. The trend in ionization energy as you go across a period is INCREASING. Therefore, the element on the far left (K) has the lowest ionization energy, and the element on the far right (Br) has the highest ionization energy.

3. Circle the element with the highest electronegativity and put a square around the element with the lowest electronegativity:

Cu K Ni Br

- a. Explain why you made these choices: All of the elements are in the same period. The trend in electronegativity as you go across a period is INCREASING. Therefore, the element on the far left (K) has the lowest electronegativity, and the element on the far right (Br) has the highest electronegativity.

4. For each of the following groups: Circle the element with the largest atomic radius and put a square around the element with the smallest atomic radius:

5. a. O C Be Ne Same Period

- b. Na Rb Fr H Same Group

- c. Pb C Sn Si Same Group

- d. Au W S Fr Ne Zn Challenge

6. For each of the following groups: Circle the element with the highest ionization energy and put a square around the element with the lowest ionization energy:

- a. O C Be Ne Same Period

- b. Na Rb Fr H Same Group

- c. Pb C Sn Si Same Group

- d. Au W S Fr Ne Zn Challenge

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Scott C. Dulebohn



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Table of Contents Periodic Trends Extension Questions Worksheet Answers

1. Understanding the eBook Periodic Trends Extension Questions Worksheet Answers
 - The Rise of Digital Reading Periodic Trends Extension Questions Worksheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Extension Questions Worksheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Trends Extension Questions Worksheet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Extension Questions Worksheet Answers
 - Personalized Recommendations
 - Periodic Trends Extension Questions Worksheet Answers User Reviews and Ratings
 - Periodic Trends Extension Questions Worksheet Answers and Bestseller Lists
5. Accessing Periodic Trends Extension Questions Worksheet Answers Free and Paid eBooks
 - Periodic Trends Extension Questions Worksheet Answers Public Domain eBooks
 - Periodic Trends Extension Questions Worksheet Answers eBook Subscription Services
 - Periodic Trends Extension Questions Worksheet Answers Budget-Friendly Options

6. Navigating Periodic Trends Extension Questions Worksheet Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Trends Extension Questions Worksheet Answers Compatibility with Devices
 - Periodic Trends Extension Questions Worksheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Trends Extension Questions Worksheet Answers
 - Highlighting and Note-Taking Periodic Trends Extension Questions Worksheet Answers
 - Interactive Elements Periodic Trends Extension Questions Worksheet Answers
8. Staying Engaged with Periodic Trends Extension Questions Worksheet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends Extension Questions Worksheet Answers
9. Balancing eBooks and Physical Books Periodic Trends Extension Questions Worksheet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends Extension Questions Worksheet Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodic Trends Extension Questions Worksheet Answers
 - Setting Reading Goals Periodic Trends Extension Questions Worksheet Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Trends Extension Questions Worksheet Answers
 - Fact-Checking eBook Content of Periodic Trends Extension Questions Worksheet Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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