

1. Circle the element with the largest atomic radius and put a square around the element with the smallest atomic radius:

Cu   **K**   Ni   **Br**

- a. Explain why you made these choices: *All of the elements are in the same period. The trend in atomic radius as you go across a period is DECREASING. Therefore, the element on the far left (K) is the largest, and the element on the far right (Br) is the smallest.*

2. Circle the element with the highest ionization energy and put a square around the element with the lowest ionization energy:

Cu   **K**   Ni   **Br**

- a. Explain why you made these choices: *All of the elements are in the same period. The trend in ionization energy as you go across a period is INCREASING. Therefore, the element on the far left (K) has the lowest ionization energy, and the element on the far right (Br) has the highest ionization energy.*

3. Circle the element with the highest electronegativity and put a square around the element with the lowest electronegativity:

Cu   **K**   Ni   **Br**

- a. Explain why you made these choices: *All of the elements are in the same period. The trend in electronegativity as you go across a period is INCREASING. Therefore, the element on the far left (K) has the lowest electronegativity, and the element on the far right (Br) has the highest electronegativity.*

4. For each of the following groups: Circle the element with the largest atomic radius and put a square around the element with the smallest atomic radius:

- 5.
- O   C   **Be**   **Ne**   Same Period
  - Na   Rb   **Fr**   **H**   Same Group
  - Pb**   **C**   Sn   Si   Same Group
  - Au   W   S   **Fr**   **Ne**   Zn   Challenge

6. For each of the following groups: Circle the element with the highest ionization energy and put a square around the element with the lowest ionization energy:

- O   C   **Be**   **Ne**   Same Period
- Na   Rb   **Fr**   **H**   Same Group
- Pb**   **C**   Sn   Si   Same Group
- Au   W   S   **Fr**   **Ne**   Zn   Challenge

# Periodic Table Trends Test Answer Key

**Alyn G. McFarland, James Napier, Roy  
White**



## **Periodic Table Trends Test Answer Key:**

## **Periodic Table Trends Test Answer Key** Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has are more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Periodic Table Trends Test Answer Key**," written by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we will delve in to the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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