

# The Periodic Law

## SECTION 2

Select the correct answer to the following questions for the given properties.

1. \_\_\_\_\_ is the number of electrons in the outermost shell of an atom.
  - (a) something on the periodic table
  - (b) something on Mendeleev's original table
  - (c) something on Mendeleev's original table
  - (d) equal to the atomic number
2. \_\_\_\_\_ Mendeleev stated that elements with similar chemical properties and chemical properties of elements appeared at regular intervals in his periodic table.
  - (a) always
  - (b) sometimes
  - (c) always
  - (d) sometimes
3. \_\_\_\_\_ The modern periodic law states that
  - (a) all elements with the same properties are listed in the same group
  - (b) the physical and chemical properties of elements are periodic functions of their atomic number
  - (c) elements with similar properties of their properties are listed
  - (d) the chemical properties of elements are periodic functions of their atomic number
4. \_\_\_\_\_ The discovery of the noble gases changed Mendeleev's periodic table.
  - (a) adding a new
  - (b) period
  - (c) series
  - (d) group
5. \_\_\_\_\_ The noble gases have properties of the noble gases is that they are
  - (a) inert
  - (b) unreactive
  - (c) highly reactive
  - (d) highly reactive
6. \_\_\_\_\_ Lanthanum, the first element in Group 3, has an atomic number of 57. How many elements in the group have an atomic number of
  - (a) 4
  - (b) 5
  - (c) 6
  - (d) 7
7. How many elements have an atomic number of 78 and are elements of 10?
  - (a) 1
  - (b) 2
  - (c) 3
  - (d) 4
8. How many elements are in the group?
  - (a) 1
  - (b) 2
  - (c) 3
  - (d) 4
9. How many elements are in the group?
  - (a) 1
  - (b) 2
  - (c) 3
  - (d) 4

# Periodic Law Chapter 5 Test Answers

**Richard Bailey**

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author's name.

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