

## Chapter 1: Limiting Reagents and Percent Yield (Pages 1-4)

**Remember:** Limiting reagent is the reactant that will completely react and leave another reactant leftover. **Remember:** Percent yield is the amount of product that is actually produced divided by the theoretical yield, multiplied by 100.

**Remember:** Limiting reagent is the reactant that will completely react and leave another reactant leftover. **Remember:** Percent yield is the amount of product that is actually produced divided by the theoretical yield, multiplied by 100.



The reaction is balanced. The limiting reagent is the reactant that will completely react and leave another reactant leftover.

2H <sub>2</sub>	20.0 g
O <sub>2</sub>	16.0 g
H <sub>2</sub> O	36.0 g
H <sub>2</sub>	20.0 g

1. 20.0 g of H<sub>2</sub> and 16.0 g of O<sub>2</sub> are mixed. The limiting reagent is H<sub>2</sub>. The theoretical yield of H<sub>2</sub>O is 36.0 g.



2. 20.0 g of H<sub>2</sub> and 16.0 g of O<sub>2</sub> are mixed. The limiting reagent is H<sub>2</sub>. The theoretical yield of H<sub>2</sub>O is 36.0 g.

3. 20.0 g of H<sub>2</sub> and 16.0 g of O<sub>2</sub> are mixed. The limiting reagent is H<sub>2</sub>. The theoretical yield of H<sub>2</sub>O is 36.0 g.



4. 20.0 g of H<sub>2</sub> and 16.0 g of O<sub>2</sub> are mixed. The limiting reagent is H<sub>2</sub>. The theoretical yield of H<sub>2</sub>O is 36.0 g.



# Percent Yield And Limiting Reactants Quiz

**LM Reder**



## Percent Yield And Limiting Reactants Quiz:

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past SQA exam questions Answers section allows students to check their understanding of the standards required for Higher level success

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