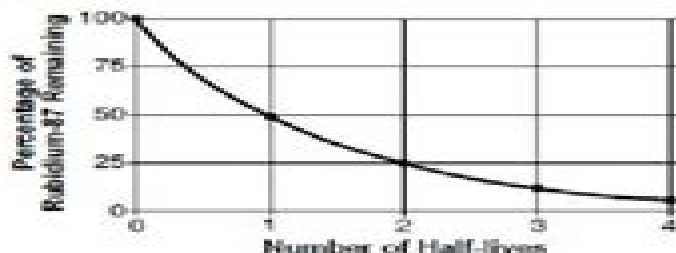


Name: _____

Answer all questions on the answer document. This is a TAKE HOME TEST due on Wednesday, June 7.

1. The graph below shows the radioactive decay of rubidium-87.



What percentage of rubidium-87 atoms will be left after four half-lives?

- A. 25.0%
 B. 12.5%
 C. 6.25%
 D. 3.125%
2. Given the equation representing a reaction at equilibrium:
- $$\text{H}_2\text{S}(\text{aq}) + \text{CH}_3\text{NH}_2(\text{aq}) \rightleftharpoons \text{HS}^-(\text{aq}) + \text{CH}_3\text{NH}_3^+(\text{aq})$$
- According to one acid-base theory, the forward reaction is classified as an acid-base reaction because
- A. H_2S is a H^+ donor and CH_3NH_2 is a H^+ acceptor
 B. CH_3NH_2 is a H^+ donor and H_2S is a H^+ acceptor
 C. HS^- and CH_3NH_3^+ are both H^+ donors
 D. CH_3NH_3^+ and HS^- are both H^+ acceptors
3. Which ionic equation represents a spontaneous reaction that can occur in a voltaic cell?
- A. $\text{Cu}(\text{s}) + \text{Zn}(\text{s}) \rightarrow \text{Cu}^{2+}(\text{aq}) + \text{Zn}^{2+}(\text{aq})$
 B. $\text{Cu}(\text{s}) + \text{Zn}^{2+}(\text{aq}) \rightarrow \text{Cu}^{2+}(\text{aq}) + \text{Zn}(\text{s})$
 C. $\text{Cu}^{2+}(\text{aq}) + \text{Zn}(\text{s}) \rightarrow \text{Cu}(\text{s}) + \text{Zn}^{2+}(\text{aq})$
 D. $\text{Cu}^{2+}(\text{aq}) + \text{Zn}^{2+}(\text{aq}) \rightarrow \text{Cu}(\text{s}) + \text{Zn}(\text{s})$

4. Given the equation representing a reaction:
- $$\text{Sn}^{4+}(\text{aq}) + 2\text{e}^- \rightarrow \text{Sn}^{2+}(\text{aq})$$

Which term best describes this reaction?

- A. ionization
 B. neutralization
 C. oxidation
 D. Reduction

5. The table below shows the atomic mass and natural abundance of the two naturally occurring isotopes of lithium.

Naturally Occurring Isotopes of Lithium

Isotope	Atomic Mass (u)	Natural Abundance (%)
Li-6	6.015	7.6
Li-7	7.016	92.4

Which numerical setup can be used to determine the atomic mass of naturally occurring lithium?

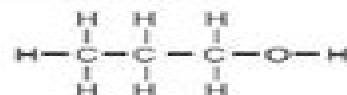
- A. $(7.6)(6.015 \text{ u}) + (92.4)(7.016 \text{ u})$
 B. $(0.076)(6.015 \text{ u}) + (0.924)(7.016 \text{ u})$

$$\frac{(7.6)(6.015 \text{ u}) + (92.4)(7.016 \text{ u})}{2}$$

 C.
$$\frac{(0.076)(6.015 \text{ u}) + (0.924)(7.016 \text{ u})}{2}$$

 D.
$$\frac{(7.6)(6.015 \text{ u}) + (92.4)(7.016 \text{ u})}{2}$$

6. Given the formulas representing two compounds at standard pressure:



1-propanol



methyl ethyl ether

The compounds can be differentiated by their

- A. boiling points
 B. gram-formula masses
 C. numbers of hydrogen atoms
 D. percent compositions by mass of carbon
7. Given the formula representing a compound:



What is the IUPAC name of this compound?

- A. 2-chloroheptane
 B. 6-chloroheptane
 C. 2,2-dichloroheptane
 D. 6,6-dichloroheptane

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 - The Rise of Digital Reading Regents Chemistry The Periodic Law Worksheet
 - Advantages of eBooks Over Traditional Books
2. Identifying Regents Chemistry The Periodic Law Worksheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Regents Chemistry The Periodic Law Worksheet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Regents Chemistry The Periodic Law Worksheet
 - Personalized Recommendations
 - Regents Chemistry The Periodic Law Worksheet User Reviews and Ratings
 - Regents Chemistry The Periodic Law Worksheet and Bestseller Lists
5. Accessing Regents Chemistry The Periodic Law Worksheet Free and Paid eBooks
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 - Regents Chemistry The Periodic Law Worksheet eBook Subscription Services
 - Regents Chemistry The Periodic Law Worksheet Budget-Friendly Options

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 - Regents Chemistry The Periodic Law Worksheet Enhanced eBook Features
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 - Highlighting and Note-Taking Regents Chemistry The Periodic Law Worksheet
 - Interactive Elements Regents Chemistry The Periodic Law Worksheet
8. Staying Engaged with Regents Chemistry The Periodic Law Worksheet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Regents Chemistry The Periodic Law Worksheet
9. Balancing eBooks and Physical Books Regents Chemistry The Periodic Law Worksheet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Regents Chemistry The Periodic Law Worksheet
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Regents Chemistry The Periodic Law Worksheet
 - Setting Reading Goals Regents Chemistry The Periodic Law Worksheet
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Regents Chemistry The Periodic Law Worksheet
 - Fact-Checking eBook Content of Regents Chemistry The Periodic Law Worksheet
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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