

Chapter 25 - Nuclear Chemistry

The study of _____ and their uses in chemistry

Radiation -

- The penetrating rays & particles emitted by a radioactive source

Radioactivity -

- Process that occurs when nuclei *change* _____

Radioisotopes -

- Atoms containing _____
- Recall... What are isotopes?

○

How do nuclear reactions differ from chemical reactions?

Chemical Reactions

- Atoms become stable by _____ or _____ electrons
- Rate _____ changed by catalysts, temperature, pressure, etc.
- Mass & charge _____

Nuclear Reactions

- Nuclei become stable by undergoing changes that _____ of energy
- Rate _____ be changed
- Charge conserved, mass _____
- A _____ small amount of mass is converted into energy

Nuclear Stability

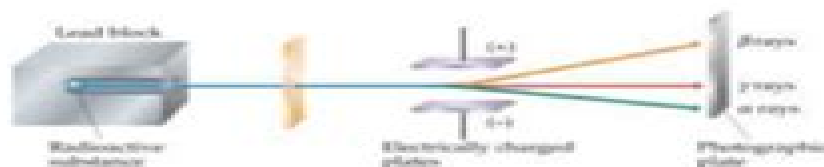
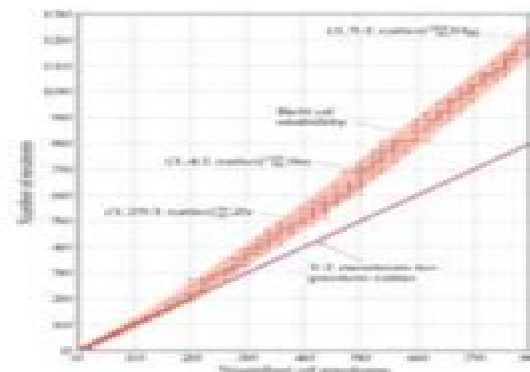
- Most nuclei are stable
- The stability of a nucleus depends on its neutron to proton _____
- The stable nuclei are in the ' _____ '

Radioactive Decay

- Occurs in atoms with too _____ or too _____ neutrons
- The nucleus will attempt to become more stable by releasing energy through _____
- Transmutation -
 - when the unstable nucleus of one element is _____ into a stable nucleus of a *different* element
 - Occurs naturally and artificially

Types of Radioactive Decay

- Alpha particles,
- Beta particles,
- Positron emission
- Gamma rays,



Regents Chemistry Nuclear Chemistry Review

M Walker



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