

Periodic Trends

1. Discuss the importance of Mendeleev's periodic law.

No

2. Identify each element as a metal, metalloid or nonmetal.

(a) fluorine	<u>non-metal</u>
(b) germanium	<u>metalloid</u>
(c) zinc	<u>metal</u>
(d) phosphorous	<u>non-metal</u>
(e) lithium	<u>metal</u>

3. Give two examples of elements for each category:

(a) noble gases	<u>Helium and Argon</u>
(b) halogens	<u>Bromine and Chlorine</u>
(c) alkali metals	<u>Sodium and Lithium</u>
(d) alkaline earth metals	<u>Barium and Strontium</u>

4. What trend in atomic radius do you see as you go down a group/family on the periodic table? What causes this trend?

As you go down a group/family on a periodic table the atomic radius increases, and the number of shells and energy levels are added. That is why there is a greater distance between the nucleus and the outermost orbital.

5. What trend in atomic radius do you see as you go across a period/row on the periodic table? What causes this trend?

When you go across a period/row on a periodic table the nuclear charge increases and attraction between nucleus and electrons gets stronger which leads the atomic radius to decrease.

6. Circle the atom in each pair that has the largest atomic radius.

(a) **Al** B
(b) **S** O

(c) **Br** Cl
(d) **Na** Al

(e) **O** F
(f) **Mg** **Ca**

7. Define ionization energy (IE.)

Is the energy to remove an electron.

8. Is it easier to form a positive ion with an element that has a high ionization energy or with an element that has a low ionization energy? Explain.

It is easier to form a positive ion with an element that has a low ionization energy, because it's easier to remove an electron which makes an ion.

9. Use the concept of ionization energy (IE) to explain why sodium forms a 1+ ion, Na^+ , but magnesium forms a 2+ ion, Mg^{2+} .

Sodium only needs to lose 1 electron which is why it forms a 1+ ion. While magnesium needs to lose 2 electron and has a higher 1st ionization energy.

Regents Chem The Periodic Law Worksheet

M Walker



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