#### Assessment 17 Worksheet: Redox Reactions & Oxidation Numbers

Complete the following questions, and upload the file to the dropbox.

 Why can an oxidation reaction not happen without a reduction reaction? Your response should clearly communicate your understanding of the nature of redox reactions.
 (1 point)

An oxidation reaction can not happen without a reduction reaction because for an oxidation reaction there must also be a reduction happening at the same time. This is because oxidation happens because of the "compound" that causes the oxidation, this "compound" must be reduced otherwise oxidation will not occur.

2. Determine the oxidation number of the specified element in each of the following.

(6 points)
Nitrogen in NF<sub>3</sub> +3

Sulfur in Se 0

Sulfur in S8 0

Chromium in CrO 2 +5\_\_\_\_

Phosphorus in P2O5 \_+5

Carbon in C12H22O11 0

Carbon in CHCl3 +2

3. Determine the oxidation number for each element in the following substances. (6 points)

\* H<sub>2</sub>SO<sub>3</sub> hydrogen: <u>+1</u> sulfur: <u>+4</u> oxygen: <u>-2</u>

\* OH oxygen: <u>-2</u> hydrogen: <u>+1</u>

\* HPO<sub>4</sub> hydrogen: <u>+1</u> phosphorus: +5 oxygen: <u>-2</u>

Determine whether each reaction is a redox reaction by finding the oxidation numbers.
 Show the oxidation numbers to justify your answer. (6 points)

\* H<sub>2</sub>O<sub>2</sub> + 2 Fe(OH)<sub>2</sub> [] 2 Fe(OH)<sub>3</sub>
H<sub>2</sub>O<sub>2</sub> + 2 Fe(OH)<sub>2</sub> [] 2 Fe(OH)<sub>3</sub>

\*\* PCl<sub>3</sub> + 3 H<sub>2</sub>O [] H<sub>3</sub>PO<sub>3</sub> + 3 HCl

\*\* PCl<sub>3</sub> + 3 H<sub>2</sub>O [] H<sub>3</sub>PO<sub>3</sub> + 3 HCl

Redox or Not? Redox

Redox or Not? Not a Redox

# **Redox Reactions Chapter Assessment Answers**

**B** Lingard

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