

CHAPTER

08

REDOX REACTION

- + The chemical reaction which contain oxidation as well as reduction half reaction is called redox reaction.
- + An atom or ion or molecule which undergoes reduction (gain of electrons) is called oxidant (while an atom or molecule which undergoes oxidation) (loss of electrons) is called reductant.
- + Oxygen shows positive oxidation states (+1 and +2) only in OF_2 and O_2F_2 .
- + Maximum oxidation state of an element never more than its group number.
- + a) Oxidation state of Mn in $KMnO_4$ is +7.
b) Oxidation state of Cr in $K_2Cr_2O_7$ is +6.
c) Oxidation state of Cr in CrO_3 is +6.
- + a) Oxidation state of 'N' in N_3H is -1/3.
b) Oxidation state of 'N' in NH_4NO_3 (NH_4^+ and NO_3^-) are -3 and +5.
- + a) Oxidation state of 'Fe' in $FeSO_4$ is +2.
b) Oxidation state of 'Fe' in Fe_3O_4 is +8/3.
- + In $xMnO_4^- + yH_2O + zI^- \rightarrow 2MnO_2 + 8OH^- + 3I_2$
 $x = 2, y = 4, z = 6$
- + In reaction of $P_4 + xH_2O + yOH^- \rightarrow PH_3 + zH_2PO_2^-$
 $x = 3, y = 3, z = 3$ (i.e. $x = y = z$)
- + $CH_4 + 2O_2 \rightarrow CO_2 + 2H_2O$ is an example for chemical combustion.
- + In all combustion reactions oxygen undergoes reduction (0 to -2).
- + In decomposition reaction chemical compound split into two or more simpler substances.
 $[2H_2O \rightarrow 2H_2 + O_2]$
- + In decomposition of $MgCO_3$, no changes in oxidation numbers. So it is not a redox reaction
 $[MgCO_3 \rightarrow MgO + CO_2]$
- + $Zn_{(s)} + CuSO_{4(aq)} \rightarrow ZnSO_{4(aq)} + Cu_{(s)}$ is a displacement reaction.
- + In disproportionation reaction the same element undergoes both oxidation and reduction simultaneously.
- + When Br_2 and I_2 dissolve in CCl_4 solvent to give reddish brown and purple colours respectively. So, they are detected by extracting into CCl_4 . This is also called "Layer Test".
- + Comproportion is a reverse of disproportionation reaction

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Brendan G. Carr



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