

Redox practice worksheet

Name: _____

Date: _____

- In which substance is the oxidation number of nitrogen zero?
A. NH_3 B. N_2 C. NO_2 D. N_2O
- What is the oxidation number of carbon in NaHCO_3 ?
A. +6 B. +2 C. -4 D. +4
- In the reaction $\text{Al}^0 + \text{Cr}^{3+} \rightarrow \text{Al}^{3+} + \text{Cr}^0$, the reducing agent is
A. Al^0 B. Cr^{3+} C. Al^{3+} D. Cr^0

- In the reaction $2\text{K} + \text{Cl}_2 \rightarrow 2\text{KCl}$, the species oxidized is
A. Cl_2 B. Cl^- C. K D. K^+
- As an S^{2-} ion is oxidized to an S^0 atom, the number of protons in its nucleus
A. decreases B. increases
C. remains the same

- Given the probable reaction for the nickel-cadmium battery:
$$2\text{Ni}(\text{OH})_2 + \text{Cd} \rightleftharpoons 2\text{Ni}(\text{OH})_2 + \text{Cd}(\text{OH})_2$$

Which species is oxidized during the discharge of the battery?

A. Ni^{2+} B. Ni^{2+} C. Cd^0 D. Cd^{2+}

- If element X forms the oxides XO and X_2O_3 , the oxidation numbers of element X are
A. +1 and +2 B. +2 and +3
C. +1 and +3 D. +2 and +4
- Oxygen has a positive oxidation number in the compound
A. H_2O B. H_2O_2 C. OF_2 D. IO_2
- What is the oxidation number of sulfur in H_2SO_4 ?
A. 0 B. -2 C. +6 D. +4

- In the equation $\text{Cu(s)} + 2\text{Ag}^+(\text{aq}) \rightarrow \text{Cu}^{2+}(\text{aq}) + 2\text{Ag(s)}$, the oxidizing agent is
A. Cu^0 B. Ag^+ C. Cu^{2+} D. Ag^0
- In the reaction $\text{Al} + \text{Cr}^{3+} \rightarrow \text{Al}^{3+} + \text{Cr}$, the reducing agent is
A. Al B. Cr^{3+} C. Al^{3+} D. Cr
- In the compound Na_2HPO_4 , which element has a negative oxidation number?
A. H B. O C. P D. Na

Redox Reactions Answers Holt Chemistry

L Darling-Hammond



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