



# **basic education**

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**REPUBLIC OF SOUTH AFRICA**

**SENIOR CERTIFICATE/  
NATIONAL SENIOR CERTIFICATE**

**GRADE 12**

**PHYSICAL SCIENCES: PHYSICS (P1)**

**NOVEMBER 2020**

**MARKS: 150**

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# Physical Science Exemplar Grade 12 2014

**Scott C. Dulebohn**

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web a write a balanced equation for the reaction  $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$  b determine the theoretical yield of  $\text{Al}_2\text{O}_3$  160.0 g  $\text{Al}$   $\times \frac{1 \text{ mol Al}}{27 \text{ g Al}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} \times 101.96 \text{ g/mol} = 160.0 \text{ g}$  c determine the percent yield  $\frac{101.96 \text{ g}}{160.0 \text{ g}} \times 100 = 63.7\%$

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