



1. The complete photoelectron spectrum of an unknown element is given above.
  - (a) Draw an X above the peak that corresponds to the orbital with electrons that are, on average, closest to the nucleus. Justify your answer in terms of Coulomb's law.
  - (b) Based on the spectrum, write the complete electron configuration of the element.
  - (c) On the graph, draw the peak(s) corresponding to the valence electrons of the element that has one more proton in its nucleus than the unknown element has.

# Released Ap Chemistry Exams

**RJ Alexander**



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