

Relative Mass and the Mole

How can atoms be counted using a balance?

Why?

Consider the following equation for a chemical reaction: $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$

This can be interpreted as two molecules of hydrogen and one molecule of oxygen combining to form two water molecules. But how often do chemists limit their reactions to one or two molecules? Usually a reaction is done with an unimaginable number of molecules. How then do chemists know they have the right mix? The molecules need to be quickly counted! How do we count molecules? The answer is the unit called the **mole**. This activity will start by considering two egg farmers (a chicken farmer and a quail farmer). They produce such large numbers of eggs that they can't count them all individually, so they count in dozens of eggs in some cases, while in other cases they use mass. Weighing is often easier than counting!

Model 1 – Eggs

Chicken		Quail		Ratio of numbers of eggs	Ratio of masses of eggs
Number of eggs in the sample	Mass of the sample	Number of eggs in the sample	Mass of the sample		
1	37.44 g	1	2.34 g	1 : 1	16 : 1
10	374.4g	10	23.4g	1:1	16:1
438	16394	438	1.023g	1:1	16:1
1 dozen	449.3g	1 dozen	28.08g	1:1	16:1
1 million	37.	1 million	2.340,000 g	1:1	16:1

- Consider the data in Model 1.
 - What is the mass of a standard chicken egg?
37.44g
 - What is the mass of a standard quail egg?
2.34g
 - Show mathematically how the 16:1 ratio of masses was calculated in the last column of Model 1.

$$\frac{37.44}{2.34} = 16$$
- Use a calculator to complete the table in Model 1. Divide the work among group members. Reduce all ratios to the lowest whole numbers possible.

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KJ Lindholm-Leary



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