

Experiment 10: Acid-Base Equilibria

How does the pH of a solution change when an acid is added?

Objectives

1. Calculate the pH of a solution of a weak acid and its conjugate base.
 2. Calculate the pH of a solution of a weak acid and its conjugate base when the acid is partially neutralized.
 3. Calculate the pH of a solution of a weak acid and its conjugate base when the acid is completely neutralized.
 4. Calculate the pH of a solution of a weak acid and its conjugate base when the acid is partially neutralized and the solution is buffered.
 5. Calculate the pH of a solution of a weak acid and its conjugate base when the acid is completely neutralized and the solution is buffered.
 6. Calculate the pH of a solution of a weak acid and its conjugate base when the acid is partially neutralized and the solution is buffered and the pH is measured.

Materials and Equipment

Solutions		Reagents		Volume of Indicator	Volume of Indicator
Volume of 0.1 M Acetic Acid	Volume of 0.1 M Sodium Acetate	Volume of 0.1 M Acetic Acid	Volume of 0.1 M Sodium Acetate	0.1 mL	0.1 mL
10 mL	0 mL	10 mL	0 mL	0.1 mL	0.1 mL
5 mL	5 mL	5 mL	5 mL	0.1 mL	0.1 mL
0 mL	10 mL	0 mL	10 mL	0.1 mL	0.1 mL
10 mL	0 mL	10 mL	0 mL	0.1 mL	0.1 mL
10 mL	0 mL	10 mL	0 mL	0.1 mL	0.1 mL

Procedure

1. Prepare the solutions as specified in the table above.

2. Measure the pH of each solution.

3. Record the pH values in the table below.

4. Calculate the pH of each solution.

Discussion

1. How does the pH of a solution of a weak acid change when its conjugate base is added? Explain.

Relative Mass The Mole Pogil Key

Tao Wei



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