- 1. Identify the attractive forces associated with...
 - a. Ionic Bonds:

Attractive electrostatic forces between negative charged ion and adjacent positively charged ion

b. Covalent Bonds:

Electrostatic forces between negatively charged electrons and adjacent positively charged nucleus

- 2. Identify two similarities between ionic and covalent bonds
- Both involve valence electron clouds
- Both involve electrostatic attractions between oppositely charged species
- The formation of bond bonds begins with the valence electrons of two atoms experiencing the attractive force of adjacent positive nuclei
- Both Ionic and Covalent bonds are strong
- Identify two difference between ionic and covalent bonds
- Ionic Bonds involve the transfer of valance electrons forming ions whereas covalent bonds involve the sharing of valence electrons with no ion formation
- Ionic bonds form only between metals and non-metals whereas covalent bonds usually (but not always) form between two non-metals
- Ionic bonding does not result in molecule formation whereas covalent bonding usually does
- The attractive forces associated with covalent bonds are the electrostatic forces between negatively charged electrons and adjacent positively charged nuclei
- Complete the following table by writing formulas of the compounds formed from the pairs of elements.

Elements	Compound Formula	ΔEN Value	Nature of Bonds Present
(a) rubidium and oxygen	Rb ₂ O	2.7	ionic
(b) strontium and bromine	SrBr ₂	1.8	lonic
(c) carbon and sulphur	CS ₂	0	covalent
(d) silicon and chlorine	SiCl ₄	1.2	polar covalent

5. Magnesium is a metal and sulphur is a non-metal. Compare the AEN value for these elements in MgS to AEN value for the non-metal hydrogen and oxygen in water (H₂O). Which of the two compounds possesses a greater amount of ionic character in its bond?

 Δ EN for MgS = 1.3 Δ EN for H₂0 = 1.4

We see that the bonds in water actually possess slightly more ionic character than those in magnesium sulphide even though the former compound contacts two non-metals and the latter compound contains a metal and a non-metal

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