

Chem 12 Electrochemistry : Quiz 2 - 120

Part 1

1) A particular reaction has a predicted E° of -2.25 V. This indicates that the reaction would :

- a) occur rapidly b) produce energy c) go to completion
d) be non-spontaneous

2) Corrosion of a metal means that the metal :

- a) acts as a cathode b) undergoes oxidation
c) undergoes reduction d) acts as an oxidizing agent

3) Zinc metal is prepared by electrolysis as follows:



What is the minimum voltage necessary for the electrolysis reaction to proceed ?

- a) 2.12 b) 1.58 c) 1.96 d) 3.48

4) Which metal would provide the best cathodic protection for iron pipes ?

- a) Ag b) Sn c) Mg d) Cu

5) Which half-reaction best describes what is taking place at the anode during the electrolysis of a 1.0 M NaCl solution ?

- a) $\text{Na}^{+} + \text{e}^{-} \rightarrow \text{Na}$ b) $2\text{Cl}^{-} \rightarrow \text{Cl}_2(\text{g}) + 2\text{e}^{-}$
c) $2\text{H}^{+}(10^{-7}\text{M}) + 2\text{e}^{-} \rightarrow \text{H}_2(\text{g})$ d) $\text{H}_2\text{O} \rightarrow 1/2 \text{O}_2(\text{g}) + 2\text{H}^{+}(10^{-7}\text{M}) + 2\text{e}^{-}$

6) Which one of the following converts chemical energy to electrical energy ?

- a) Industrial production of aluminum b) Nickel plating
c) Discharging a "dry" cell d) Production of Na(s)

7) Which is the unit for the energy per coulomb of electrons ?

Redox And Electrochemistry Quiz Answers

Yijin Wang

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the author's name.

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Table of Contents Redox And Electrochemistry Quiz Answers

1. Understanding the eBook Redox And Electrochemistry Quiz Answers
 - The Rise of Digital Reading Redox And Electrochemistry Quiz Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Redox And Electrochemistry Quiz Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Redox And Electrochemistry Quiz Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Redox And Electrochemistry Quiz Answers
 - Personalized Recommendations
 - Redox And Electrochemistry Quiz Answers User Reviews and Ratings
 - Redox And Electrochemistry Quiz Answers and Bestseller Lists
5. Accessing Redox And Electrochemistry Quiz Answers Free and Paid eBooks
 - Redox And Electrochemistry Quiz Answers Public Domain eBooks
 - Redox And Electrochemistry Quiz Answers eBook Subscription Services

- Redox And Electrochemistry Quiz Answers Budget-Friendly Options
- 6. Navigating Redox And Electrochemistry Quiz Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Redox And Electrochemistry Quiz Answers Compatibility with Devices
 - Redox And Electrochemistry Quiz Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Redox And Electrochemistry Quiz Answers
 - Highlighting and Note-Taking Redox And Electrochemistry Quiz Answers
 - Interactive Elements Redox And Electrochemistry Quiz Answers
- 8. Staying Engaged with Redox And Electrochemistry Quiz Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Redox And Electrochemistry Quiz Answers
- 9. Balancing eBooks and Physical Books Redox And Electrochemistry Quiz Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Redox And Electrochemistry Quiz Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Redox And Electrochemistry Quiz Answers
 - Setting Reading Goals Redox And Electrochemistry Quiz Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Redox And Electrochemistry Quiz Answers
 - Fact-Checking eBook Content of Redox And Electrochemistry Quiz Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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