

Name: \_\_\_\_\_

Class: \_\_\_\_\_

## OXIDATION NUMBERS WORKSHEET

The *oxidation number* of an atom is the apparent charge assigned to it in a particular molecule, ion or compound. Certain rules are followed in assigning oxidation numbers.

Use rules at the right to assign oxidation numbers to each element in each of the given formulas.

Example

	OXIDATION NO.	RULES USED
$\text{H}_2\text{O}$	<u>H=+1, O= -2</u>	<u>6 &amp; 7</u>
$\text{N}_2$	<u>N= 0</u>	<u>1</u>

**Rules** The oxidation number of:

1. A free element in the uncombined state is 0.
2. A monatomic ion equals the charge on the ion.
3. Oxygen in a compound is generally -2
4. Hydrogen in a compound is +1
5. In a neutral compound the sum of the oxidation numbers for all the atoms is 0.
6. In a polyatomic ion the sum of the oxidation numbers for all the atoms is equals the charge on the ion.
7. Group 1 metals is +1.
8. Group 2 metals is +2.

		RULES USED
1.	$\text{Cl}_2$	_____
2.	$\text{Cl}^-$	_____
3.	$\text{Na}$	_____
4.	$\text{Na}^+$	_____
5.	$\text{KCl}$	_____
6.	$\text{H}_2\text{S}$	_____
7.	$\text{CaO}$	_____
8.	$\text{H}_2\text{SO}_4$	_____
9.	$\text{NO}_3^-$	_____
10.	$\text{Cr}_2\text{O}_7^{2-}$	_____
11.	$\text{NH}_4\text{Cl}$	_____
12.	$\text{NH}_3$	_____
13.	$\text{NO}_2$	_____
14.	$\text{CaH}_2$ (calcium hydride)	_____
15.	$\text{Na}_2\text{O}_2$ (sodium peroxide)	_____

# Redox Ws 2 Oxidation Numbers Answers

**Carson-Dellosa Publishing**



## Redox Ws 2 Oxidation Numbers Answers:

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*Metalloenzymes* Claudiu T. Supuran, William Alexander Donald, 2023-08-30

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Indian Journal of Chemistry ,2008 Fiscal Year 1986 Department of Energy Authorization (basic Research Programs) United States. Congress. House. Committee on Science and Technology. Subcommittee on Energy Development and Applications,1985

**Chemistry Homework** Frank Schaffer Publications, Joan DiStasio, 1996-03 Includes the periodic table writing formulas balancing equations stoichiometry problems and more

*Conversion of Water and CO<sub>2</sub> to Fuels using Solar Energy* Oomman K. Varghese, Flavio L. Souza, 2024-03-26

*Conversion of Water and CO<sub>2</sub> to Fuels using Solar Energy* Comprehensive Resource for Understanding the Emerging Solar Technologies for Hydrogen Generation via Water Splitting and Carbon based Fuel Production via CO<sub>2</sub> Recycling Fossil fuel burning is the primary source of carbon in the atmosphere The realization that such burning can harm the life on our planet has led to a surge in research activities that focus on the development of alternative strategies for energy conversion Fuel generation using solar energy is one of the most promising approaches that has received widespread attention The fuels produced using sunlight are commonly referred to as solar fuels This book provides researchers interested in solar fuel generation a comprehensive understanding of the emerging solar technologies for hydrogen generation via water splitting and carbon based fuel production via CO<sub>2</sub> recycling The book presents the fundamental science technologies techno economic analysis and most importantly the materials that are being explored to establish artificial methods of fuel

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 field by presenting the core concepts basic principles and newest advances      **Dissertation Abstracts International** ,2009

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 such a central role in the formation of sedimentary sulfides It will be of interest to scientists in a number of disciplines  
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