

# Redox Exercise (Part 2)

## Chapter 20 Worksheet: Redox ANSWERS

1. Determine what is oxidized and what is reduced in each reaction. Identify the oxidizing agent and the reducing agent, also.



11. Give the oxidation number of each kind of atom or ion.

a. sulfate	b. Sn	c. $\text{S}^{2-}$	d. $\text{Fe}^{3+}$	e. $\text{Sn}^{4+}$	f. nitrate	g. ammonium
2-	0	2-	3+	4+	1-	1+

12. Calculate the oxidation number of chromium in each of the following.

a. $\text{Cr}_2\text{O}_3$	b. $\text{Na}_2\text{Cr}_2\text{O}_7$	c. $\text{CrSO}_4$	d. chromate	e. dichromate
3+	6+	2+	6+	6+

13. Use the changes in oxidation numbers to determine which elements are oxidized and which are reduced in these reactions. (Note: it is not necessary to use balanced equations)



14. For each reaction in problem 13, identify the oxidizing agent and reducing agent.

- |                               |                          |
|-------------------------------|--------------------------|
| a. oxidizing agent: sulfur    | reducing agent: carbon   |
| b. oxidizing agent: nitrogen  | reducing agent: iodine   |
| c. oxidizing agent: manganese | reducing agent: chloride |
| d. oxidizing agent: nitrogen  | reducing agent: antimony |

# Redox Reactions Chemistry If8766 Answers 9

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## **Table of Contents Redox Reactions Chemistry If8766 Answers 9**

1. Understanding the eBook Redox Reactions Chemistry If8766 Answers 9
  - The Rise of Digital Reading Redox Reactions Chemistry If8766 Answers 9
  - Advantages of eBooks Over Traditional Books
2. Identifying Redox Reactions Chemistry If8766 Answers 9
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Redox Reactions Chemistry If8766 Answers 9
  - User-Friendly Interface
4. Exploring eBook Recommendations from Redox Reactions Chemistry If8766 Answers 9
  - Personalized Recommendations
  - Redox Reactions Chemistry If8766 Answers 9 User Reviews and Ratings
  - Redox Reactions Chemistry If8766 Answers 9 and Bestseller Lists

5. Accessing Redox Reactions Chemistry If8766 Answers 9 Free and Paid eBooks
  - Redox Reactions Chemistry If8766 Answers 9 Public Domain eBooks
  - Redox Reactions Chemistry If8766 Answers 9 eBook Subscription Services
  - Redox Reactions Chemistry If8766 Answers 9 Budget-Friendly Options
6. Navigating Redox Reactions Chemistry If8766 Answers 9 eBook Formats
  - ePub, PDF, MOBI, and More
  - Redox Reactions Chemistry If8766 Answers 9 Compatibility with Devices
  - Redox Reactions Chemistry If8766 Answers 9 Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Redox Reactions Chemistry If8766 Answers 9
  - Highlighting and Note-Taking Redox Reactions Chemistry If8766 Answers 9
  - Interactive Elements Redox Reactions Chemistry If8766 Answers 9
8. Staying Engaged with Redox Reactions Chemistry If8766 Answers 9
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Redox Reactions Chemistry If8766 Answers 9
9. Balancing eBooks and Physical Books Redox Reactions Chemistry If8766 Answers 9
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Redox Reactions Chemistry If8766 Answers 9
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Redox Reactions Chemistry If8766 Answers 9
  - Setting Reading Goals Redox Reactions Chemistry If8766 Answers 9
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Redox Reactions Chemistry If8766 Answers 9
  - Fact-Checking eBook Content of Redox Reactions Chemistry If8766 Answers 9
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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