Balance each redox reaction in acid solution using either method (show the work).

$$8 \quad H_3O_3 + Cr_2O_7^3 \rightarrow O_3 + Cr^{3*}$$

9.
$$TeO_3^2 + N_2O_4 \rightarrow Te + NO_3^2$$

State of the change that represents oxidation, reduction or neither. Use oxidation #s
Remember that if the oxidation #increases it means oxidation and when it decreases
it mean reduction!

18.
$$MnO_2 \rightarrow Mn_2O_3$$

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Barron's Educational Series, Neil D. Jespersen, Pamela Kerrigan

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