

Name \_\_\_\_\_

# Physical Science Chemical Bonding Study Guide

## A. Covalent & Ionic Compounds/Bonds

1. Which ion accepts electrons, and which from it? (anion/cation) \_\_\_\_\_  
which cation it gains and electron it loses from it? \_\_\_\_\_ (anion/cation)
2. How placed in the halogens group has \_\_\_\_\_ valence electrons, and number is \_\_\_\_\_ 1 smaller in size is. This is because of the \_\_\_\_\_ rules.
3. How \_\_\_\_\_ bond in the addition between oppositely charged ions. (ionic covalent bond) is \_\_\_\_\_
4. On a covalent bond, atoms \_\_\_\_\_, electrons, shared by all. Ionic bond, atoms \_\_\_\_\_, electrons.
5. (Electron shell) Which side from left shared completely is a covalent bond. It is \_\_\_\_\_, and which side from shared equally is the \_\_\_\_\_ bond. (partial) is a covalent. In ionic, gain electrons, or negative charged by, losing losing the \_\_\_\_\_
6. How atoms can have up to \_\_\_\_\_ valence electrons. This usually happens in the periodic table that all have 8 valence electrons, or called \_\_\_\_\_
7. Atoms can gain, lose, or share \_\_\_\_\_ by becoming more stable.

## B. Formulas for Ionic Compounds – Put in the right order

Write correct chemical formula	Aluminum	Fluorine	Phosphorus	Chlorine
EXPLANATION IS				

## C. Write the chemical formulas for each of the following ions: compound is

Write your own ionic compounds, and use arrows to show the transfer of electrons. Use your periodic table to find their oxidation numbers. They will combine so that their oxidation numbers cancel out to zero. Use the criss cross method if their oxidation numbers do not cancel.

# Physical Science Chemical Bonding Study Guide

**MOHAN, RADHA**



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- Chapter 5: Conclusion

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