

# PHYSICAL CHEMISTRY

## (Complete Syllabus)

### SECTION - A

#### Straight Objective Type

This section contains 220 multiple choice questions. Each question has 4 choices (1), (2), (3) and (4), out of which **ONLY ONE** is correct.

---

Choose the correct answer :

- The mass of oxygen required to produce enough CO which completely reduces  $x$  gm of  $\text{Fe}_2\text{O}_3$ . Suppose molecular weight of  $\text{O}_2 = y$ , molecular weight of  $\text{Fe}_2\text{O}_3 = z$   
(1)  $\frac{2xz}{3y}$  (2)  $\frac{3xy}{2z}$   
(3)  $\frac{2y}{3xz}$  (4)  $\frac{3z}{2xy}$
- The number of lone pair of electrons present in 2.56 gm of orthorhombic sulphur will be  
(1)  $0.02 N_0$  (2)  $0.16 N_0$   
(3)  $0.32 N_0$  (4)  $0.08 N_0$
- A 10.0 g of sample of mixture of calcium chloride and sodium chloride is treated with  $\text{Na}_2\text{CO}_3$  to precipitate the calcium as calcium carbonate. This  $\text{CaCO}_3$  is heated to convert all calcium to CaO and final mass of CaO is 1.62 gms. The % by mass of  $\text{CaCl}_2$  in the original mixture is  
(1) 15.2%  
(2) 32.1%  
(3) 21.8%  
(4) 11.07%
- 1000 ml of oxygen at STP were passed through an ozoniser and the resulting volume was 888 ml at STP. This quantity of ozonised oxygen is passed through excess of KI solution. The weight of  $\text{I}_2$  liberated is  
(1) 25.4 g (2) 12.7 g  
(3) 2.54 g (4) 1.27 g
- $x_1$  ml solution having normality 1 is diluted to get a solution of normality  $y_2$ . What would be volume required for dilution?  
(1)  $x_1 \left(1 - \frac{y_1}{y_2}\right)$  (2)  $x_1 \left(\frac{y_2}{y_1} - 1\right)$   
(3)  $x_1 \left(\frac{y_1}{y_2} - 1\right)$  (4)  $x_1 \left(1 + \frac{y_1}{y_2}\right)$
- 1.625 g of an interhalogen compound  $\text{ICl}_x$  was subjected to electric discharge which split it into constituents and brought to STP. It occupied 112 ml. The interhalogen compound is  
(1)  $\text{ICl}$  (2)  $\text{ICl}_3$   
(3)  $\text{ICl}_5$  (4)  $\text{ICl}_7$
- If 12.25 g of a sample of  $\text{KClO}_3$  of 75% purity and rest vapourisable impurities is strongly heated to give  $\text{KCl}$ ,  $\text{KClO}_4$  and 0.32 gm  $\text{O}_2$ . The mass of residue obtained is  
(1) 1.769 g (2) 7.096 g  
(3) 5.327 g (4) 8.85 g
- A 2 g sample of xenon reacts with fluorine. The mass of compound produced is 3.158 g. The empirical formula of the compound is  
(1)  $\text{XeF}_2$  (2)  $\text{XeF}_4$   
(3)  $\text{XeF}_6$  (4)  $\text{XeF}_8$
- 1 g of carbonate of metal was dissolved in 25 ml of 1 N HCl. The resulting liquid required 5 ml of 1 N NaOH for neutralisation. The equivalent weight

# Physical Chemistry Exam Questions

**Richard H. Langley**



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## **Table of Contents Physical Chemistry Exam Questions**

1. Understanding the eBook Physical Chemistry Exam Questions
  - The Rise of Digital Reading Physical Chemistry Exam Questions
  - Advantages of eBooks Over Traditional Books
2. Identifying Physical Chemistry Exam Questions
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Physical Chemistry Exam Questions
  - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Chemistry Exam Questions
  - Personalized Recommendations
  - Physical Chemistry Exam Questions User Reviews and Ratings
  - Physical Chemistry Exam Questions and Bestseller Lists
5. Accessing Physical Chemistry Exam Questions Free and Paid eBooks
  - Physical Chemistry Exam Questions Public Domain eBooks
  - Physical Chemistry Exam Questions eBook Subscription Services
  - Physical Chemistry Exam Questions Budget-Friendly Options
6. Navigating Physical Chemistry Exam Questions eBook Formats

- ePub, PDF, MOBI, and More
  - Physical Chemistry Exam Questions Compatibility with Devices
  - Physical Chemistry Exam Questions Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Physical Chemistry Exam Questions
    - Highlighting and Note-Taking Physical Chemistry Exam Questions
    - Interactive Elements Physical Chemistry Exam Questions
  8. Staying Engaged with Physical Chemistry Exam Questions
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Physical Chemistry Exam Questions
  9. Balancing eBooks and Physical Books Physical Chemistry Exam Questions
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Physical Chemistry Exam Questions
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Physical Chemistry Exam Questions
    - Setting Reading Goals Physical Chemistry Exam Questions
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Physical Chemistry Exam Questions
    - Fact-Checking eBook Content of Physical Chemistry Exam Questions
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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